

Acids and Bases Worksheet 2

1. In the reaction: $\text{CO}_3^{2-} + \text{H}_2\text{O} \rightleftharpoons \text{HCO}_3^- + \text{OH}^-$
the carbonate ion is acting as a(n):
(A) Arrhenius base
(B) Arrhenius acid
(C) Bronsted-Lowry base
(D) Bronsted-Lowry acid
2. Which of the following represents a Bronsted-Lowry conjugate acid-base pair?
(A) SO_3^{2-} and SO_2 (B) CO_3^{2-} and CO (C) H_3O^+ and H_2 (D) NH_4^+ and NH_3
3. Which of the following would not make a good buffering system?
(A) SO_4^{2-} and H_2SO_4
(B) HCO_3^- and H_2CO_3
(C) NH_3 and NH_4^+
(D) CH_3COO^- and CH_3COOH
4. A water solution of which of the following compounds will turn blue litmus red
(A) K_2CO_3 (B) NH_4Cl (C) NaOH (D) NaCl
5. 10.0 ml of a solution of HCl required 12.5 ml of 0.400 M $\text{Ba}(\text{OH})_2$ for complete neutralization. How many moles of HCl were present in the sample?
(A) 5.00×10^{-3} (B) 1.00×10^{-2} (C) 1.00 (D) 2.00
6. If 1.00 L of 1.00 M CH_3COOH is mixed with 0.25 mole of solid NaOH (assume no volume change), what will be the pH of the resulting solution? ($K_a = 1.8 \times 10^{-5}$).
(A) 4.14 (B) 4.27 (C) 4.74 (D) 5.35
7. What is the molarity of CH_3COOH in vinegar containing 4.0% CH_3COOH by mass and having a density of 1.02 g/ml?
(A) 0.50 M (B) 0.68 M (C) 0.75 M (D) 1.36 M
8. If 5.00 ml of 15.4 M HNO_3 is diluted to 250 ml, what is the pH of the resulting solution?
(A) 0.51 (B) 0.76 (C) 1.45 (D) 2.89

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Sir Norman Lockyer



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Chemistry McGraw-Hill Staff,2001-03 **Organic Chemistry** Seyhan N. Ege,1998-11 *Fundamentals of Environmental Chemistry, Third Edition* Stanley E. Manahan,2011-03-05 Written by an expert using the same approach that made the previous two editions so successful *Fundamentals of Environmental Chemistry Third Edition* expands the scope of book to include the strongly emerging areas broadly described as sustainability science and technology including green chemistry and industrial ecology The new edition includes Increased emphasis on the applied aspects of environmental chemistry Hot topics such as global warming and biomass energy Integration of green chemistry and sustainability concepts throughout the text More and updated questions and answers including some that require Internet research Lecturers Pack on CD ROM with solutions manual PowerPoint presentations and chapter figures available upon qualifying course adoptions The book provides a basic course in chemical science including the fundamentals of organic chemistry and biochemistry The author uses real life examples from environmental chemistry green chemistry and related areas while maintaining brevity and simplicity in his explanation of concepts Building on this foundation the book covers environmental chemistry broadly defined to include sustainability aspects green chemistry industrial ecology and related areas These chapters are organized around the five environmental spheres the hydrosphere atmosphere geosphere biosphere and the anthrosphere The last two chapters discuss analytical chemistry and its relevance to environmental chemistry Manahan s clear concise and readable style makes the information accessible regardless of the readers level of chemistry knowledge He demystifies the material for those who need the basics of chemical science for their trade profession or study curriculum as well as for readers who want to have an understanding of the fundamentals of sustainable chemistry in its crucial role in maintaining a livable planet

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[Solving Equilibrium Problems with Applications to Qualitative Analysis](#) Steven S. Zumdahl, 1989

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[Advanced Cardiac Life Support](#) Bruce R. Shade, 1986

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