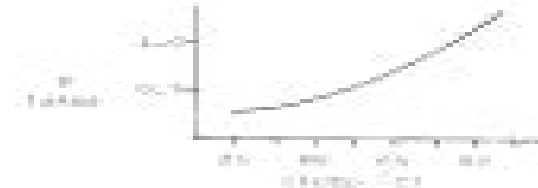


- This is practice. Do NOT reveal yourself of finding out what you are capable of doing. Be sure you follow the testing conditions outlined below.
- Do NOT USE a CALCULATOR. You may use ONE of the four periodic tables.
- Try to work out as many of T-2 into per questions. Time yourself. It is important that you practice working for speed.
- When when time is up, continue working until finished as necessary.

- Surface tension in a liquid is due to the fact that
 - intermolecular forces on the surface molecules all pulling toward the interior
 - liquids need kinetic lowest energy
 - potential energy is increased for molecules on the surface
 - surface molecules are attracted in all directions
 - all of the above
- In which one of the following will dipole-dipole attractions play the most significant role in the intermolecular attractions?
 - HCl
 - H_2O
 - NaCl
 - NH_3
 - Br_2
- In which type of substances are London dispersion forces the only type of IMF?
 - polar molecules
 - metals
 - nonpolar compounds
 - non-polar molecules
 - ionic compounds
- The heat of vaporization of H_2O , at its boiling point ($+100^\circ\text{C}$) is 18.8 kJ/mol . What mass of H_2O can be vaporized at its boiling point with 100 kJ of energy?
 - $\frac{(100)(0.1)}{(18.8)}$
 - $\frac{(18.8)(100)}{(0.1)}$
 - $(0.1)(18.8)(100)(0.1)$
 - $\frac{(18.8)(0.1)}{(100)}$
 - $\frac{(100)(100)}{(18.8)}$
- Which one of the following substances exhibits the strongest intermolecular forces of attraction?
 - CH_4
 - C_2H_6
 - CH_3OH
 - CH_3CN
 - C_2H_5
- For which substance would you predict the highest heat of vaporization?
 - F_2
 - H_2O
 - HBr
 - NaCl
 - HCl
- Which of the following will change the equilibrium vapor pressure of a liquid?
 - Heat up or cool down the liquid
 - Increase the volume of the container
 - Change the air pressure above the liquid
 - I only
 - I and II only
 - I and III only
 - II and III only
 - I, II, and III
- Which substance below exhibits the weakest IMF?
 - H_2O
 - SO_2
 - CO_2
 - SiO_2
 - HCl
- Which one of the following is linked with a type of IMF that the substance does not display?
 - $\text{NaCl}_{(s)}$ ion-dipole
 - CCl_4 dipole-dipole
 - H_2O hydrogen bonding
 - C_2H_6 covalent bonding
 - HCl ionic
- At what temperature will the liquid (whose vapor pressure is shown below) boil if the air pressure is reduced to 380 mmHg ?
 
 - 100°C
 - 50°C
 - 75°C
 - 100°C
 - the liquid will not boil at this pressure

Blue Devil Chem Mc Test

Katrin Zwirgmaier



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