

AP* Chemistry GASES

The gaseous state of matter is the simplest and best-understood state of matter. You inhale approximately 8,500 L of air each day. This amounts to about 25 lbs of air. Breathing is a three-step process: inhaling, gas exchange with the circulatory system, and exhaling. Approximately 80% pressure differences created by your body allow you to breathe. Nearly the Earth's entire atmosphere is made up of only five gases: nitrogen, oxygen, water vapor, argon, and carbon dioxide.

PROPERTIES OF GASES

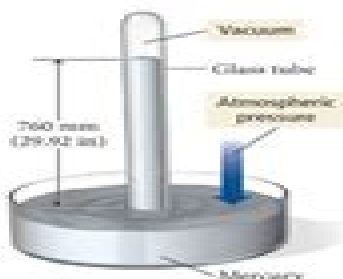
Only **FOUR** quantities are needed to define the state of a gas:

1. the quantity of the gas, n (in moles)
2. the temperature of the gas, T (in **KELVINS**)
3. the volume of the gas, V (in liters)
4. the pressure of the gas, P (in atmospheres)

A gas uniformly fills any container, is easily compressed & mixes completely with any other gas.



The Mercury Barometer



Pressure and Density



GAS PRESSURE is a measure of the force that a gas exerts on its container. Force is the physical quantity that interferes with inertia. Gravity is the force responsible for weight.

Force = mass $\times$ acceleration; (Newton's 2nd Law)
The SI units follow:

$$N = kg \times m/s^2$$

Pressure— Force/ unit area; N/m^2 ; which is the definition of 1.0 Pascal

Barometer—invented by Evangelista Torricelli in 1643; uses the height of a column of mercury to measure gas pressure (especially atmospheric P)

$$1 \text{ mm of Hg} = 1 \text{ torr}$$

$$1.00 \text{ atm} = 760.00 \text{ mm Hg} = 760.00 \text{ torr} = 29.92 \text{ in Hg} = 14.7 \text{ psi} = 101.325 \text{ kPa} \approx 10^5 \text{ Pa}$$

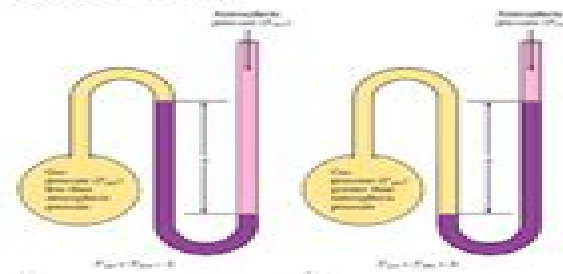
At sea level all of the above define **STANDARD PRESSURE**.

The SI unit of pressure is the Pascal (named after Blaise Pascal);

$$1 \text{ Pa} = 1 \text{ N} / \text{m}^2$$

The **manometer**—a device for measuring the pressure of a gas in a container. The pressure of the gas is given by h [the difference in mercury levels] in units of torr (equivalent to mm Hg).

- a) Gas pressure = atmospheric pressure – h
- b) Gas pressure = atmospheric pressure + h



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