

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)
- Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	<u>5.640 M</u>	Potassium chromate	<u>3.350 M</u>
Cobalt chloride	<u>4.330 M</u>	Nickel (II) chloride	<u>5.210 M</u>
Potassium dichromate	<u>0.510 M</u>	Copper sulfate	<u>1.380 M</u>
Gold (III) chloride	<u>2.30 M</u>	Potassium permanganate	<u>0.480 M</u>

Concentration And Molarity Phet Chemistry Labs

Robert R. Redfield



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Table of Contents Concentration And Molarity Phet Chemistry Labs

1. Understanding the eBook Concentration And Molarity Phet Chemistry Labs
 - The Rise of Digital Reading Concentration And Molarity Phet Chemistry Labs
 - Advantages of eBooks Over Traditional Books
2. Identifying Concentration And Molarity Phet Chemistry Labs
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concentration And Molarity Phet Chemistry Labs
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concentration And Molarity Phet Chemistry Labs
 - Personalized Recommendations
 - Concentration And Molarity Phet Chemistry Labs User Reviews and Ratings
 - Concentration And Molarity Phet Chemistry Labs and Bestseller Lists
5. Accessing Concentration And Molarity Phet Chemistry Labs Free and Paid eBooks
 - Concentration And Molarity Phet Chemistry Labs Public Domain eBooks
 - Concentration And Molarity Phet Chemistry Labs eBook Subscription Services

- Concentration And Molarity Phet Chemistry Labs Budget-Friendly Options
- 6. Navigating Concentration And Molarity Phet Chemistry Labs eBook Formats
 - ePub, PDF, MOBI, and More
 - Concentration And Molarity Phet Chemistry Labs Compatibility with Devices
 - Concentration And Molarity Phet Chemistry Labs Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concentration And Molarity Phet Chemistry Labs
 - Highlighting and Note-Taking Concentration And Molarity Phet Chemistry Labs
 - Interactive Elements Concentration And Molarity Phet Chemistry Labs
- 8. Staying Engaged with Concentration And Molarity Phet Chemistry Labs
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concentration And Molarity Phet Chemistry Labs
- 9. Balancing eBooks and Physical Books Concentration And Molarity Phet Chemistry Labs
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concentration And Molarity Phet Chemistry Labs
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Concentration And Molarity Phet Chemistry Labs
 - Setting Reading Goals Concentration And Molarity Phet Chemistry Labs
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concentration And Molarity Phet Chemistry Labs
 - Fact-Checking eBook Content of Concentration And Molarity Phet Chemistry Labs
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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