

Concept Review

Section 8-1: Covalent Bonds

Answer the following items in the space provided.

1. What does a molecular bond-forming reaction look like?
2. Why is the H_2 molecule more stable than two separate hydrogen atoms?
3. Explain why the stability difference in H_2 does not show that bond-forming is more energetically favorable.
4. What does a molecular bond-forming reaction look like?
5. What is enthalpy change?
6. Describe potential energy changes when atoms form hydrogen bonds. Explain why the reaction is exothermic.
7. In terms of energy, why is the H_2 bond stable?
8. Why is bond length an average value? How is bond length?
9. Describe the physical properties of molecules that are formed by covalent bonds.

Explain the use of structural formula by showing a water molecule. The following is the structural formula for water:

structure	structural formula	electron	electron
structural formula	electron	electron	electron
1. Bond _____ is the average distance between two bonded atoms in the molecule.	electron	electron	electron
2. Bond _____ is the average distance between two bonded atoms in the molecule.	electron	electron	electron
3. The average of all the average bond lengths between two atoms is the average bond length.	electron	electron	electron
4. In _____, the average of all the average bond lengths between two atoms is the average bond length.	electron	electron	electron
5. In _____, the average of all the average bond lengths between two atoms is the average bond length.	electron	electron	electron

Concept Review Section Covalent Bonds 6 Answers

**Eldon Enger, Frederick Ross, David
Bailey**



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