

Name: \_\_\_\_\_

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### Gas Laws Test Review

Select the correct answer for the following questions. For any calculation you **MUST** show ALL work with units. For any multiple choice question you **MUST** list the textbook page number or section that help you select the correct answer.

- If temperature and pressure are held constant, the volume and number of moles of a gas are:  
a. directly proportional      b. independent of each other  
c. not enough information given      d. equal      e. inversely proportional
- If volume and number of moles of a gas are held constant, the temperature and pressure are:  
a. directly proportional      b. independent of each other  
c. not enough information given      d. equal      e. inversely proportional
- If temperature and number of moles of a gas are held constant, the volume and pressure are:  
a. directly proportional      b. independent of each other  
c. not enough information given      d. equal      e. inversely proportional
- If pressure and number of moles of a gas are held constant, the volume and temperature are:  
a. directly proportional      b. independent of each other  
c. not enough information given      d. equal      e. inversely proportional
- One mole of  $\text{NH}_3$  at STP will have what volume?  
a. 22.4 L      b. 44g      c. 1.0 L      d. 24.5 L
- Which of the following statements is true about the kinetic molecular theory?  
a) Pressure is due to the collisions of the gas particles with the walls of the container.  
b) Adding an ideal gas to a closed container will cause an increase in temperature.  
c) The volume of a gas particle is considered to be small about 0.1 mL.  
d) Gas particles have random motion.  
e) At least two of these statements are true.
- Which of the following statements is true concerning ideal gases?  
a) The gas particles in a sample exert attraction for one another.  
b) At STP 1.0L of Ar gas contains about twice the number of atoms as 1.0L of Ne gas.  
c) A gas exerts pressure as a result of the collisions of the gas molecules with the walls of the container.  
d) The temperature of the gas sample is indirectly related to the average velocity of the gas particles.
- Use the kinetic molecular theory of gases to predict what would happen to a closed sample of a gas whose temperature increased while its volume decreased.  
a) Its pressure would decrease      c) Its pressure would hold constant.  
b) Its pressure would increase.      d) the number of moles would increase.
- Perform the following conversions of pressure units:  $5.0 \times 10^5 \text{ Pa} = \underline{\hspace{2cm}} \text{ atm}$   
a.  $4.9 \times 10^4$       b.  $1.7 \times 10^5$       c.  $4.3 \times 10^4$       d.  $2.5 \times 10^4$       e.  $9.8 \times 10^4$
- The pressure of a given mass of gas varies directly with the temperature when the volume remains constant.  
a. Boyle's Law      b. combined gas law      c. Gay-Lussac's Law  
d. ideal gas law      e. Charles's Law
- If the temperature is held constant, what will happen to the pressure exerted by a gas if the volume of the container is decreased?  
a. increase      b. decrease      c. remain the same      d. increase or decrease (it depends)
- If the temperature of a sample of gas is decreased at constant pressure, the volume of the gas will:  
a. increase      b. decrease      c. remain the same      d. you cannot predict this.

# Concept Review Section The Gas Laws Answer

**Victor M. Corman**



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