

# CHEMICAL EQUILIBRIUM



## EQUILIBRIUM CONSTANT 'K'.

For a General Reaction



The equilibrium constant expression is

$$K_c = \frac{[P]^p [Q]^q}{[A]^a [B]^b}$$

where  $K_c$  is the **Equilibrium Constant** (or  $K_p$  if they are all gases)

## RELATION BETWEEN $K_p$ AND $K_c$

For the Reaction



$$K_p = \frac{[P_c]^c \times [P_d]^d}{[P_a]^a \times [P_b]^b} = \frac{[C]^c [D]^d}{[A]^a [B]^b} \frac{(RT)^{c+d}}{(RT)^{a+b}}$$

$$K_p = K_c (RT)^{\Delta n_g}$$

$$\text{if } \Delta n_g = 0 \implies K_p = K_c$$

Where,  $\Delta n_g = (c+d) - (a+b)$

= no. of moles of gaseous products - no. of moles of gaseous Reactants

## WHAT DOES THE VALUE OF 'K' MEAN?

$$\text{(a) } K \gg 1$$

Reactants



Products

if  $K \gg 1$ , the reaction is product-favoured; product predominates at equilibrium.

$$\text{(a) } K \ll 1$$

Reactants



Products

if  $K \ll 1$ , the reaction is reactant-favoured; reactant predominates at equilibrium.

$$\text{(a) } K \approx 1$$

the reaction lies in the middle (no. of reactants and products)

## MAGNITUDE OF 'K'

Small ( $K < 10^{-3}$ )



Intermediate ( $10^{-3} < K < 10^3$ )



Large ( $K > 10^3$ )



# Concept Review Continued Chemical Equilibrium

**S Marginson**



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