

Name _____ Class _____ Date _____

Skills Worksheet

Concept Review**Section: Calculating Quantities in Reactions**

Complete each statement below by writing the correct term or phrase.

1. All stoichiometric calculations involving equations use _____ ratios.
2. When solving stoichiometric problems, you must _____ the equation first.
3. Balanced equations give the _____ numbers of moles of substances.
4. _____ in chemical equations provide mole ratios that can be used as conversion factors.
5. The conversion factor for converting between mass and moles is the _____ of the substance.
6. In making calculations involving _____, you must convert volume to mass.
7. To convert from volume to mass, you can use the _____ of the substance as the conversion factor.
8. When calculating the number of particles, you can use _____ as the conversion factor.

In the blanks at the left, write the letter of the choice that best completes the statement or answers the question. Consider the following problem when answering:

What mass of sulfuric acid is required to neutralize 2.65 g of potassium hydroxide? The products of the reaction are potassium sulfate and water.

- _____ 9. What should you do first after reading the problem carefully?
- a. Estimate the answer.
 - b. Calculate the molar mass of sulfuric acid.
 - c. Write a balanced chemical equation.
 - d. Convert all masses to moles.
- _____ 10. What should you do before setting up the problem?
- a. Determine the densities.
 - b. Calculate molar masses.
 - c. Convert all masses to moles.
 - d. Estimate the answer.

Concept Review Section 3 Chemistry

Y Pai

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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<u>Engineering Practice</u> ,1976	<i>Concepts in Biology</i> Eldon Enger,Frederick Ross,David Bailey,2008-01-16 Enger Ross Bailey		

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Table of Contents Concept Review Section 3 Chemistry

1. Understanding the eBook Concept Review Section 3 Chemistry
 - The Rise of Digital Reading Concept Review Section 3 Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Section 3 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concept Review Section 3 Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Section 3 Chemistry
 - Personalized Recommendations

- Concept Review Section 3 Chemistry User Reviews and Ratings
- Concept Review Section 3 Chemistry and Bestseller Lists
- 5. Accessing Concept Review Section 3 Chemistry Free and Paid eBooks
 - Concept Review Section 3 Chemistry Public Domain eBooks
 - Concept Review Section 3 Chemistry eBook Subscription Services
 - Concept Review Section 3 Chemistry Budget-Friendly Options
- 6. Navigating Concept Review Section 3 Chemistry eBook Formats
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 - Concept Review Section 3 Chemistry Compatibility with Devices
 - Concept Review Section 3 Chemistry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concept Review Section 3 Chemistry
 - Highlighting and Note-Taking Concept Review Section 3 Chemistry
 - Interactive Elements Concept Review Section 3 Chemistry
- 8. Staying Engaged with Concept Review Section 3 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concept Review Section 3 Chemistry
- 9. Balancing eBooks and Physical Books Concept Review Section 3 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concept Review Section 3 Chemistry
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Concept Review Section 3 Chemistry
 - Setting Reading Goals Concept Review Section 3 Chemistry
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concept Review Section 3 Chemistry
 - Fact-Checking eBook Content of Concept Review Section 3 Chemistry

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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