

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are $^{32}_{16}\text{S}$ and $^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 4.1
2. sub atomic 4.2
3. negative 4.3
4. protons 4.2
5. neutrons 4.2
6. nucleus 4.3
7. Rutherford 4.3
8. positive 4.3
9. electrons 4.4
10. protons (elec) 4.4
11. number 4.4
12. electrons 4.4
13. protons (elec) 4.5
14. neutrons 4.6
15. isotopes 4.6
16. atomic mass 4.7
17. 16 4.5
18. 16 4.6
19. 16 4.6
20. 18 4.6

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IM Harris



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environmental crisis and offers ideas on how we as a global community can create a sustainable future.
Focus on physical science Charles H. Heimler, 1989
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Pasqualetti, Jennifer Keahey, 2022-09-29
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Sociology Jon M. Shepard, 1999 This comprehensive mainstream text presents the basic theoretical conceptual and methodological tenets of sociology in a readable and personal manner Its balanced presentation challenges students to pose their own questions about society and to think critically about the world in which they live It presents the study of sociology as a means by which students can transcend personal and societal prejudices This Seventh Edition continues its tradition of integrating innovative pedagogy directly into the book such as a built in student study guide lecture launching CNN video segments that link to text specific topics and continued use of the SQ3R survey question read recite and review learning technique

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Digest; Review of Reviews Incorporating Literary Digest , 1903

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