

## *Solutions and Solubility*

### **Dissolving Process**

1. Solvent molecules are attracted to solute molecules or ions at the surface of the particles.
2. Solvent molecules surround the solute molecule.
3. Solvent molecules carry the solute molecule into the solution.
4. The solute molecule dissolves in a “like” solvent.

### **Factors that Affect the Rate of Dissolution**

- The **rate of dissolution** describes the rate or speed at which a solid dissolves a liquid.
- It is affected by:
  - Stirring.
  - Surface area.
  - Temperature.

### **Solubility**

- **Solubility** is the amount of a solute that will dissolve in a given volume of solvent at a given temperature and pressure.
- It depends on the solute and the solvent.
  - Solubility does not equal dissolution rate.

### **Unsaturated, Saturated, and Supersaturated Solutions**

- Types of solutions:
  - **Unsaturated solution** – the concentration of solute is **less than** the maximum concentration predicted from the solute’s solubility.
  - **Saturated solution** – the concentration of solute is **equal to** the maximum concentration predicted from the solute’s solubility, the solute crystallizes at the same rate that it dissolves.
  - **Supersaturated solution** – the concentration of solute is **greater than** the maximum concentration predicted from the solute’s solubility.

### **The Effect of Temperature on Solubility**

- Increasing temperature:
  - Increases the solubility of solids and liquids.
  - Decreases the solubility of gases.

### **The Effect of Pressure on Solubility**

- Increasing pressure:
  - Increases the solubility of gases.
  - Has little effect on the solubility of solids and liquids.

### **Solubility Graphs**

- Solubility graphs:
  - Show how solubility changes with temperature.
  - Are determined experimentally.

# Concept Review Solubility And The Dissolving Process

## Holt Chemistry

**Lingjun Ying**



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