

At the conclusion of your Fischer esterification experiment, you obtained an IR spectrum of a carboxylic acid, an alcohol, and an ester.

- a. The carbonyl stretch of carboxylic acids and esters is located between $1750\text{--}1700\text{ cm}^{-1}$. Consider this statement taken from your lab manual: 'The force required to stretch or bend a bond is the function of the strength of the bond and the masses of the bonded atoms.' Explain the difference in wavenumbers between carbonyl and hydroxyl functional groups.
- b. Carboxylic acid functional groups have an O–H stretch located between $3500\text{--}2500\text{ cm}^{-1}$. Alcohols have an OH stretch located between $3650\text{--}3200\text{ cm}^{-1}$.
 - i. Draw hydrogen bonding interactions between two carboxylic acid molecules.
 - ii. Draw hydrogen bonding interactions between two alcohols.
 - iii. Consider the effects of hydrogen bonding on bond strength and use your illustrations to explain the difference in wavenumbers between the hydroxyl groups of a carboxylic acid and the hydroxyl group of an alcohol.
- c. Carboxylic acids have a carbonyl stretch located between $1725\text{--}1700\text{ cm}^{-1}$. Esters have a carbonyl stretch located between $1750\text{--}1730\text{ cm}^{-1}$. Consider the effects of hydrogen bonding, resonance, and inductive electron withdrawal on the bond strength of the carbonyl group of carboxylic acids and esters. Explain the difference in wavenumbers between the carbonyl stretch of a carboxylic acid and that of an ester.

Conclusion For Esterification Experiment

Johns Hopkins University



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