

# COVALENT BONDING

Name \_\_\_\_\_

Covalent bonding occurs when two or more nonmetals share electrons, attempting to attain a stable octet of electrons at least part of the time. For example:

*H wants 2 electrons  
all other atoms 8*



Note that hydrogen is content with 2, not 8, electrons.

*electrons in the middle share*

Show how covalent bonding occurs in each of the following pairs of atoms. Atoms may share one, two or three pairs of electrons.

*wanted available share*

<p>1. <math>\text{H} + \text{H} (\text{H}_2)</math></p> <p>W - 4</p> <p>A - 2</p> <p>S - 2</p>	<p><i>wanted 2 + 2</i></p> <p><i>available</i></p> <p><math>\text{H} \cdot \cdot \text{H}</math></p> <p><math>\text{H} \cdot \quad \cdot \text{H}</math></p>	<p><math>\text{H} - \text{H}</math></p> <p><i>single bond = share 2 e<sup>-</sup></i></p>
<p>2. <math>\text{F} + \text{F} (\text{F}_2)</math></p> <p>W - 8 + 8 = 16</p> <p>A - 7 + 7 = 14</p> <p>S - 2</p>	<p><math>\cdot \ddot{\text{F}} \cdot \cdot \ddot{\text{F}} \cdot</math></p> <p><i>2 e<sup>-</sup></i></p>	<p><math>\cdot \ddot{\text{F}} - \ddot{\text{F}} \cdot</math></p> <p><i>single bond 2 e<sup>-</sup></i></p>
<p>3. <math>\text{O} + \text{O} (\text{O}_2)</math></p> <p>W - 8 + 8 = 16</p> <p>A - 6 + 6 = 12</p> <p><i>(S = 4)</i></p>	<p><math>\cdot \ddot{\text{O}} \cdot \cdot \ddot{\text{O}} \cdot</math></p> <p><i>4 e<sup>-</sup></i></p>	<p><math>\cdot \ddot{\text{O}} = \ddot{\text{O}} \cdot</math></p> <p><i>4 double bond</i></p>
<p>4. <math>\text{N} + \text{N} (\text{N}_2)</math></p> <p>W - 8 + 8 = 16</p> <p>A - 5 + 5 = 10</p> <p>S - 6</p> <p><i>3 e<sup>-</sup></i></p>	<p><math>\cdot \ddot{\text{N}} \cdot \cdot \ddot{\text{N}} \cdot</math></p>	<p><math>\cdot \ddot{\text{N}} \equiv \ddot{\text{N}} \cdot</math></p>
<p>5. <math>\text{C} + \text{O} (\text{CO}_2)</math></p> <p>W - 8 + 8 + 8 = 24</p> <p>A - 4 + 6 + 6 = 16</p> <p>S - 8</p>	<p><math>\cdot \ddot{\text{O}} \cdot \cdot \ddot{\text{C}} \cdot \cdot \ddot{\text{O}} \cdot</math></p>	<p><math>\cdot \ddot{\text{O}} = \text{C} = \ddot{\text{O}} \cdot</math></p>
<p>6. <math>\text{H} + \text{O} (\text{H}_2\text{O})</math></p> <p>W - 2 + 2 + 8 = 12</p> <p>A - 1 + 1 + 6 = 8</p> <p>S - 4</p>	<p><math>\text{H} \cdot \cdot \ddot{\text{O}} \cdot \cdot \text{H}</math></p>	<p><math>\text{H} - \ddot{\text{O}} - \text{H}</math></p>

# Covalent Bonding Practice Problems Answers

**J Elliott**



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*Student Solutions Manual* Darrell Ebbing, 2005-03-17 Provides solutions to odd numbered Practice Problems General Problems and Cumulative Skills Problems plus answers to Review Questions

*General Organic and Biological Chemistry* Kenneth W. Raymond, 2013-01-10 General Organic and Biological Chemistry 4th Edition Binder Ready Version has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry An integrated approach is employed in which related general chemistry organic chemistry and biochemistry topics are presented in adjacent chapters This approach helps students see the strong connections that exist between these three branches of chemistry and allows instructors to discuss these interrelationships while the material is still fresh in students minds This text is an unbound binder ready edition

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