

Covalent Bonding

How does a covalent bond form?

- A **covalent bond** forms when two atoms **share** one or more pairs of electrons.
- Atoms in covalent bonds typically have **similar electronegativities** and are **usually non-metals**.
- The shared electrons allow each atom to achieve a full outer electron shell, similar to the noble gas configuration.

Types of Covalent Bonds

- **Single Bond:** One pair of electrons is shared (e.g., H_2 , Cl_2).
- **Double Bond:** Two pairs of electrons are shared (e.g., O_2 , CO_2).
- **Triple Bond:** Three pairs of electrons are shared (e.g., N_2).



Dot-and-Cross Diagrams

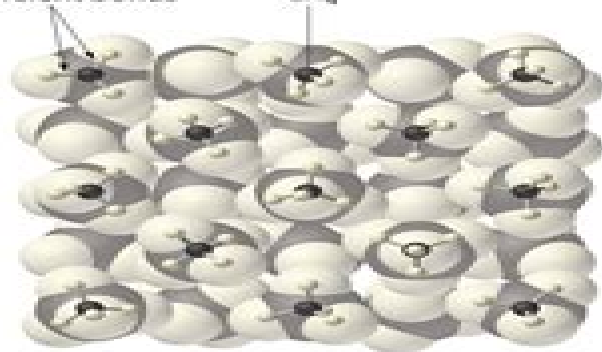
- Only the outer electron shells (valence electrons) are shown.
- Shared pairs of electrons between atoms are shown as pairs of dots and crosses.
- It represents how the atoms bond and achieve stable electron configurations.

Example: Water H_2O



Covalent bonds

CH_4



Properties of Covalent Compounds

- **Molecular Structure:** They are made up of discrete molecules with **weak intermolecular forces**.
- **Melting and Boiling Points:** Generally **lower** than ionic compounds because the forces between molecules are weaker.
- **Electrical Conductivity:** **Poor conductors** of electricity in all states, as they do not have free ions or electrons.
- **Solubility:** Often **soluble in nonpolar solvents** and may be insoluble in water, depending on their polarity.

Covalent Bonding Study Guide

**Katherine J. Denniston, Denniston
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Table of Contents Covalent Bonding Study Guide

1. Understanding the eBook Covalent Bonding Study Guide
 - The Rise of Digital Reading Covalent Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Covalent Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Study Guide
 - Personalized Recommendations
 - Covalent Bonding Study Guide User Reviews and Ratings
 - Covalent Bonding Study Guide and Bestseller Lists

5. Accessing Covalent Bonding Study Guide Free and Paid eBooks
 - Covalent Bonding Study Guide Public Domain eBooks
 - Covalent Bonding Study Guide eBook Subscription Services
 - Covalent Bonding Study Guide Budget-Friendly Options
6. Navigating Covalent Bonding Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Study Guide Compatibility with Devices
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Study Guide
 - Highlighting and Note-Taking Covalent Bonding Study Guide
 - Interactive Elements Covalent Bonding Study Guide
8. Staying Engaged with Covalent Bonding Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Study Guide
9. Balancing eBooks and Physical Books Covalent Bonding Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Covalent Bonding Study Guide
 - Setting Reading Goals Covalent Bonding Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Covalent Bonding Study Guide
 - Fact-Checking eBook Content of Covalent Bonding Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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