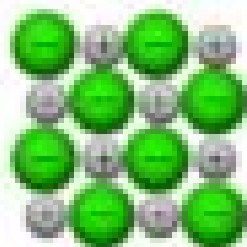


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The Chemistry

Ionic, Covalent, and metallic compounds each have physical properties that relate to their structures, so there are some tests we can do to help identify the type of bonding in a substance.

Melting Points

An ionic bond is made when ions of opposite charges come together to create a **crystal** lattice structure. In order to melt an ionic compound, the temperature has to rise until the ions move with enough force to break away from the oppositely charged ions around them.

Covalent Compounds don't break apart when they melt, so they only have to reach a temperature strong enough to pull them away from the other molecules to melt them. This is why most small covalent molecules are liquid or gas at room temperature.



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Solubility & Conductivity



Solubility and Conductivity

Ionic compounds **dissolve very easily** in water to create an aqueous solution. Since water is a polar molecule, it attracts the ions and surrounds each one as the compound dissociates in the water. Once water is full of charged particles, it can conduct electricity very easily. Tap water conducts electricity because the chemicals used to treat the water are ions, so it makes the water into a conductor.



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O García



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1. Understanding the eBook Covalent Bonds Lesson 20 Unit 4
 - The Rise of Digital Reading Covalent Bonds Lesson 20 Unit 4
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonds Lesson 20 Unit 4
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonds Lesson 20 Unit 4
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonds Lesson 20 Unit 4
 - Personalized Recommendations
 - Covalent Bonds Lesson 20 Unit 4 User Reviews and Ratings
 - Covalent Bonds Lesson 20 Unit 4 and Bestseller Lists
5. Accessing Covalent Bonds Lesson 20 Unit 4 Free and Paid eBooks
 - Covalent Bonds Lesson 20 Unit 4 Public Domain eBooks
 - Covalent Bonds Lesson 20 Unit 4 eBook Subscription Services
 - Covalent Bonds Lesson 20 Unit 4 Budget-Friendly Options
6. Navigating Covalent Bonds Lesson 20 Unit 4 eBook Formats

- ePub, PDF, MOBI, and More
- Covalent Bonds Lesson 20 Unit 4 Compatibility with Devices
- Covalent Bonds Lesson 20 Unit 4 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonds Lesson 20 Unit 4
 - Highlighting and Note-Taking Covalent Bonds Lesson 20 Unit 4
 - Interactive Elements Covalent Bonds Lesson 20 Unit 4
- 8. Staying Engaged with Covalent Bonds Lesson 20 Unit 4
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonds Lesson 20 Unit 4
- 9. Balancing eBooks and Physical Books Covalent Bonds Lesson 20 Unit 4
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonds Lesson 20 Unit 4
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonds Lesson 20 Unit 4
 - Setting Reading Goals Covalent Bonds Lesson 20 Unit 4
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonds Lesson 20 Unit 4
 - Fact-Checking eBook Content of Covalent Bonds Lesson 20 Unit 4
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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