



Covalent Compounds Unit Study Guide

Name _____
 Teacher _____
 Date: _____
 Period _____

Reference:

Chemistry in the Community, pp 181-183,
 Notes
 Naming covalent compounds worksheet
 Viscosity Lab
 Exploring Properties of Ionic and Covalent compounds

Vocabulary

Understand and apply the definitions of the following terms:

Covalent	Octet rule
Valence electrons	Polymers
e dot diagrams (Lewis dot diagrams)	Viscosity
Polar	Boiling points
Nonpolar	
Intermolecular forces	

Targeted Skills

- Diagram a covalent molecule using a e dot diagrams (Lewis dot diagrams).
- Know the following properties:
 - > Ionic compounds = the transfer of electron(s), i.e. "gain or lose"
 Metal cations & nonmetal anion (or polyatomic ion)
 Net charge = zero (criss cross charges) $X^{+n}Y^{-m} = X_nY_m$
 Dissociate (dissolve) in water
 - > Molecular compounds = sharing electrons by forming a covalent bond
 2 or more nonmetals
 Prefixes indicate # of each atom
 Polar molecules have a δ^+ & a δ^-
 Polar molecules dissolve in water
 Nonpolar molecules do not have oppositely charged ends, therefore do not dissolve in water.
 - > Octet rule = atoms share, gain, or lose e- to obtain full outer energy level (usually 8 e-)
- Name covalent compounds using prefixes: (mono, di, tri, tetra, penta, hexa, hepta, octa, nona, deca).
- Compare/contrast the arrangement of atoms in molecules to the arrangement of atoms in ionic compounds.
- Describe the influence of intermolecular forces on physical and chemical properties of covalent compounds. Example: Boiling points, viscosity, and density.
- Be able to describe properties (boiling points, melting points, viscosity, solubility, conductivity) of polymers, ionic crystals, metallic substances, atoms and molecules.

Make sure you remember from past tests?

Atomic structure	Solubility
Significant figures	Safety
Accuracy & precision	Elements
Mixtures	Ionic compound formula writing
Nomenclature	

Covalent Compounds Study Guide

Kenneth L. Henold



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