

CP Chemistry – 2nd Semester Final Review

Test Date:

Per. 6 Monday 6/22 10AM-noon

Time: 120 minutes (that's it, folks)

Basic Test Layout:

You've worked hard this year, and here is the chance to show what you know. Everything, meaning, *everything*, from this **past year** is fair game on the final exam (but the emphasis is on second semester material). The exam will be 100 multiple-choice questions that cover chapters 8-12, 18-20, 25/26 (roughly 10 questions per chapter). Some of these questions require background knowledge from the first semester. **The final is 20% of your score in this class.**

You will be allowed only a few aides on the exam: the periodic table reference sheet, a **non-graphing** calculator, unlimited scratch paper, and the packet of chapter quiz notes you have been making all semester that I will return to you in class (each page turned in has a stamp on it). *You may not add any additional pages to this packet, but you may add extra notes to existing pages. Extra pages will result lowering your final exam score by one letter grade (perhaps more depending on how severe).* You must turn in **everything** you used on the test, except your calculator, when you turn in the final.

Study Tips: Don't assume that just because you made good notes (or not) earlier in the semester, that because the final is open notes you do not need to study. You have 120 minutes to answer 100 questions – it is impossible to look every question up in your notes and finish the entire exam. Also, topics that were covered in January are most likely going to be a bit rusty. There are no "short cuts" on this test, and it will test if you really know your stuff. You have learned a lot this semester – an honest effort in studying and test taking will get you far. Don't give up the last week of the semester – Finish well!

Chapter 8 – Chemical Reactions

Balancing chemical reactions (8.1)
Four types of reactions (8.2)
Net ionic equations (8.3)
Formation of precipitates (8.3)

Chapter 9 – Stoichiometry **calculations**

Mole Road Map
Mole ratio
 $\text{mol A} \rightarrow \text{mol B}$, $\text{g A} \rightarrow \text{g B}$,
 $\text{L A} \rightarrow \text{L B}$, etc
Limiting reagent, percent yield (9.3)

Chapter 10 – States of Matter

Solids (10.3)
Liquids (10.2)
Gases (10.1) – Kinetic Theory, STP
Phase Changes/Diagram
Atmospheric pressure (relation to boiling)
Barometer
Kinetic Energy & Temperature

Chapter 11 –

Thermochemistry

Enthalpy
(endothermic/exothermic)
Heat capacity (C); heat (q)
Calories vs. calories

Calorimetry problems

($q = mCAT$)
Latent heat problems
($q = \Delta H^{\circ} \text{moles}$)
Hess's Law (11.4)

Chapter 12 – Gases

STP; variables describing a gas
Boyle's, Charles's, Gay-Lussac's laws
Gas law problems (combined gas law, ideal gas law)
Pressure (units, description, conversion)
Kinetic Energy

Chapter 18 – Solutions

Dilute vs. concentrated
Molarity (Definition and Calculations)
Concentration
%w/v, %m/v (Calculation)
Dilutions ($M_1V_1 = M_2V_2$)
solubility
factors affecting solubility

Chapter 19 – Equilibrium and Reaction Rates

Equilibrium
Le Chatelier's Principle & Problems (Shift L or R)
Equilibrium constant (Write out, not calculate)
Activation energy

Catalyst

Collision Theory
Factors affecting rates
Activation Energy diagram

Chapter 20 – Acids and Bases

Acid/Base Properties
Arrhenius, Bronstad-Lowry, Lewis definitions
pH calculations and scale
Acid and base definitions
Strong acid/base vs. weak acid/base

Chapter 25/26: Organic Chemistry

Naming
alkanes/alkenes/alkynes/alcohol
Functional groups: ether, ester, acid, amine, alcohol, etc.
Enzymes (define)

First Semester Review Topics

Mole Road Map Calculations
Naming (ion charges, polyatomic ions)
Electron Configurations
Periodic Trends
Lewis Dot Structures
Metric Conversions

Cp Chemistry Final Exam

Hussin A.Rothana



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