Chemistry

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(Chapter 4) (The d and f block Elements) (Exercise)
(Class 12)

Question 4.23:

Which metal in the first series of transition metals exhibits +1 oxidation state most frequently and why? €Answer 4.23:

In the first transition series, Cu exhibits +1 oxidation state very frequently. It is because Cu (+1) has an electronic configuration of $[Ar]3d^{10}$. The completely filled d-orbital makes it highly stable.

Question 4.24:

Calculate the number of unpaired electrons in the following gaseous ions: Mn^{3+} , Cr^{3+} , V^{3+} and Ti^{3+} . Which one of these is the more stable in aqueous solution?

__Answer 4.24:

	Gaseous ions	Number of unpaired electrons
(i)	Mn^{3+} , $[Ar]^3d^4$	4
(11)	Cr3+, [Ar]3d3	3
(1111)	V3+, [Ar]3d2	2
(iv)	Ti3+, [Ar]3d1	1

Cr3+ is the most stable in aqueous solutions owing to a t3e configuration.

Question 4.25:

Give examples and suggest reasons for the following features of the transition metal chemistry:

(i) The lowest oxide of transition metal is basic, the highest is amphoteric/acidic.

(ii) A transition metal exhibits highest oxidation state in oxides and fluorides.

(iii) The highest oxidation state is exhibited in oxoanions of a metal.

Answer 4.25:

(i) In the case of a lower oxide of a transition metal, the metal atom has a low oxidation state. This means that some of the valence electrons of the metal atom are not involved in bonding. As a result, it can donate electrons and behave as a base.

On the other hand, in the case of a higher oxide of a transition metal, the metal atom has a high oxidation state. This means that the valence electrons are involved in bonding and so, they are unavailable. There is also a high effective nuclear charge. As a result, it can accept electrons and behave as an acid.

For example, Mn^{II}O is basic and Mn^{VII}O₂ is acidic.

- (ii) Oxygen and fluorine act as strong oxidising agents because of their high electro negativities and small sizes. Hence, they bring out the highest oxidation states from the transition metals. In other words, a transition metal exhibits higher oxidation states in oxides and fluorides. For example, inOsF₆ and V₂O₅, the oxidation states of Os and V are+6 and+5 respectively.
- (iii) Oxygen is a strong oxidising agent due to its high electronegativity and small size. So, oxo-anions of a metal have the highest oxidation state. For example, in MnO₄, the oxidation state of Mn is +7.

Question 4.26:

Indicate the steps in the preparation of:

(ii) KMnO₄ from pyrolusite ore.

(i) Potassium dichromate (K₂Cr₂O₇)is prepared from chromite ore (FeCr₂O₄)in the following steps: Step (i): Preparation of sodium chromate

 $4FeCr_2O_4 + 16NaOH + 7O_2 \rightarrow 8Na_2CrO_4 + 2Fe_2O_3 + 8H_2O$

Step (ii): Conversion of sodium chromate into sodium dichromate

 $2Na_2CrO_4 + conc.H_2SO_4 \rightarrow Na_2Cr_2O_7 + NaSO_4 + H_2O$

Step (iii): Conversion of sodium dichromate to potassium dichromate

Na₂Cr₂O₇ + 2KCl → K₂Cr₂O₇ + 2NaCl

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- 1. Understanding the eBook Cumulative Review For Chemistry Chapter 4 Answers
 - The Rise of Digital Reading Cumulative Review For Chemistry Chapter 4 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Cumulative Review For Chemistry Chapter 4 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Cumulative Review For Chemistry Chapter 4 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Cumulative Review For Chemistry Chapter 4 Answers
 - Personalized Recommendations
 - Cumulative Review For Chemistry Chapter 4 Answers User Reviews and Ratings
 - Cumulative Review For Chemistry Chapter 4 Answers and Bestseller Lists
- 5. Accessing Cumulative Review For Chemistry Chapter 4 Answers Free and Paid eBooks
 - Cumulative Review For Chemistry Chapter 4 Answers Public Domain eBooks
 - Cumulative Review For Chemistry Chapter 4 Answers eBook Subscription Services
 - o Cumulative Review For Chemistry Chapter 4 Answers Budget-Friendly Options
- 6. Navigating Cumulative Review For Chemistry Chapter 4 Answers eBook Formats

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- Cumulative Review For Chemistry Chapter 4 Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Cumulative Review For Chemistry Chapter 4 Answers
 - Highlighting and Note-Taking Cumulative Review For Chemistry Chapter 4 Answers
 - Interactive Elements Cumulative Review For Chemistry Chapter 4 Answers
- 8. Staying Engaged with Cumulative Review For Chemistry Chapter 4 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Cumulative Review For Chemistry Chapter 4 Answers
- 9. Balancing eBooks and Physical Books Cumulative Review For Chemistry Chapter 4 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Cumulative Review For Chemistry Chapter 4 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Cumulative Review For Chemistry Chapter 4 Answers
 - Setting Reading Goals Cumulative Review For Chemistry Chapter 4 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Cumulative Review For Chemistry Chapter 4 Answers
 - Fact-Checking eBook Content of Cumulative Review For Chemistry Chapter 4 Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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