

Solution Concentrations REVIEW

NAME: _____

Solution → Solute (KoolAid) + Solvent (Water)

$$\text{Percent by Mass} = \frac{\text{Mass of Solute (g)}}{\text{Mass of Solution (g)}} \times 100$$

$$\text{Percent by Volume} = \frac{\text{Volume of Solute (mL)}}{\text{Volume of Solution (mL)}} \times 100$$

$$\text{Mass / Volume} = \frac{\text{Mass of solute (g)}}{\text{Volume of Solution (mL)}}$$

$$\text{Molarity (M)} = \frac{\text{Moles of solute}}{\text{Volume of Solution (L)}}$$

1. Glucose is a sugar that is found abundantly in nature. What is the percent by mass of a solution made by dissolving 163 g of glucose in 755 g of water? Do you need to know the formula of glucose? Why or why not?
2. What is the mass percent sucrose in a solution obtained by mixing 225 g of sucrose with 135 g water?
3. Determine the volume percent of toluene in a solution made by mixing 40.0 mL toluene with 75.0 mL of benzene.
4. What is the molarity of a solution prepared by dissolving 2.25g of glucose ($\text{C}_6\text{H}_{12}\text{O}_6$) in 0.500 L solution.

Concentrations Of Solutions Section Review Answers

Marco Cascella



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