

$$dG = -SdT + VdP + \sum_{i=1}^c \mu_i dN_i$$

The Gibbs Equation

$$0 = \sum_{i=1}^c \mu_i dN_i$$

When a system is in Equilibrium, the change in Gibb's energy is 0. In our example, neither the volume nor the temperature of the system change.

$$0 = \sum_{i=1}^c \mu_i dN_i = \mu_1 + \mu_2 + \dots + \mu_c$$

An expansion of the total change in chemical potential for each component with respect to the change in phase.

$$0 = \mu_1 = \mu_1^{(1)} + \mu_1^{(2)} + \dots + \mu_1^{(N_p)}$$

Because our system consists of only one component, the total chemical potential is only that contained in the component (1).

$$0 = \mu_1^{(1)} + \mu_1^{(2)}$$

Because we only have two phases, the summation of the chemical potential of those two phases is equal to 0.

$$\mu_1^{(1)} = -\mu_1^{(2)}$$

This allows us to see that at Equilibrium, the chemical potential of the liquid phase and the vapor phase are equal. The signs are opposite because on change in phase would release the Gibbs Energy that is taken in by the other during a change in phase.

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A Loxley



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Table of Contents Concept Review Chapter 1section 4 Phase Equilibrium

1. Understanding the eBook Concept Review Chapter 1section 4 Phase Equilibrium
 - The Rise of Digital Reading Concept Review Chapter 1section 4 Phase Equilibrium
 - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Chapter 1section 4 Phase Equilibrium
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concept Review Chapter 1section 4 Phase Equilibrium
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Chapter 1section 4 Phase Equilibrium
 - Personalized Recommendations
 - Concept Review Chapter 1section 4 Phase Equilibrium User Reviews and Ratings
 - Concept Review Chapter 1section 4 Phase Equilibrium and Bestseller Lists
5. Accessing Concept Review Chapter 1section 4 Phase Equilibrium Free and Paid eBooks
 - Concept Review Chapter 1section 4 Phase Equilibrium Public Domain eBooks
 - Concept Review Chapter 1section 4 Phase Equilibrium eBook Subscription Services
 - Concept Review Chapter 1section 4 Phase Equilibrium Budget-Friendly Options
6. Navigating Concept Review Chapter 1section 4 Phase Equilibrium eBook Formats

- ePub, PDF, MOBI, and More
 - Concept Review Chapter 1section 4 Phase Equilibrium Compatibility with Devices
 - Concept Review Chapter 1section 4 Phase Equilibrium Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concept Review Chapter 1section 4 Phase Equilibrium
 - Highlighting and Note-Taking Concept Review Chapter 1section 4 Phase Equilibrium
 - Interactive Elements Concept Review Chapter 1section 4 Phase Equilibrium
 8. Staying Engaged with Concept Review Chapter 1section 4 Phase Equilibrium
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concept Review Chapter 1section 4 Phase Equilibrium
 9. Balancing eBooks and Physical Books Concept Review Chapter 1section 4 Phase Equilibrium
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concept Review Chapter 1section 4 Phase Equilibrium
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Concept Review Chapter 1section 4 Phase Equilibrium
 - Setting Reading Goals Concept Review Chapter 1section 4 Phase Equilibrium
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Concept Review Chapter 1section 4 Phase Equilibrium
 - Fact-Checking eBook Content of Concept Review Chapter 1section 4 Phase Equilibrium
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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