

Concept Review

Section 8-1: Covalent Bonds

Answer the following items in the space provided.

1. What does a molecular bond-forming reaction look like?
2. Why is the H_2 molecule more stable than two separate hydrogen atoms?
3. Explain why the stability difference in H_2 does not show that bond-forming is more energetically favored.
4. What does a molecular bond-forming reaction look like?
5. What is enthalpy change?
6. Describe potential energy changes that occur when two hydrogen atoms approach each other, what forms a covalent bond?
7. In terms of energy, why is the H_2 bond stable?
8. Why is bond length an average value? How is bond energy?
9. Describe the physical properties of molecules that the covalent bond, ionic, and covalent bonds.

Complete each statement below by choosing a term from the following list. Name may be used more than once.

- | | | | |
|---------------------|---------------------|----------|----------------|
| strategy | electron-attraction | electron | ionic covalent |
| electron-attraction | electron | electron | electron |
1. Bond _____ is the force that holds atoms together. Bonding is the force that holds atoms together.
 2. Bond _____ is the force that holds atoms together. Bonding is the force that holds atoms together.
 3. The attraction of an atom for electrons is called the _____.
 4. In _____, electrons are attracted to the nucleus of the atom in which they are located.
 5. In _____, electrons are attracted to the nucleus of the atom in which they are located.

Concept Review Chapter 6 Section Covalent Bonds

Answer Key

Michael Hoelscher



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