

CHAPTER 19 REVIEW

Oxidation-Reduction Reactions**SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ All of the following should be done in the process of balancing redox equations *except*:
 - (a) adjusting coefficients to balance atoms.
 - (b) adjusting coefficients in electron equations to balance numbers of electrons lost and gained.
 - (c) adjusting subscripts to balance atoms.
 - (d) writing two separate electron equations.
2. MnCl_2 can be reduced to MnO_2 .
 - _____ a. Assign the oxidation number to Mn in these two species.
 - _____ b. How many electrons are gained per Mn atom in this reduction?
 - _____ c. If 0.50 mol of MnO_2 is reduced, how many electrons are gained?
3. Iodide ions can be oxidized to form iodine. Write the balanced oxidation half-reaction for the oxidation of iodide to iodine.

4. Some bleaches contain aqueous chlorine as the active ingredient. Aqueous chlorine is made by dissolving chlorine gas in water. Aqueous chlorine is capable of oxidizing iron(II) ions to iron(III) ions. When iron(II) ions are oxidized, chloride ions are formed.
 - a. Write equations for the two half-reactions involved. Label them oxidation or reduction.

 - b. Write the balanced ionic equation for the redox reaction between aqueous chlorine and iron(II).

 - c. Show that the equation in part b is balanced by charge.

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Table of Contents Concept Review Section Oxidation Reduction Reactions Answers

1. Understanding the eBook Concept Review Section Oxidation Reduction Reactions Answers
 - The Rise of Digital Reading Concept Review Section Oxidation Reduction Reactions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review Section Oxidation Reduction Reactions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Concept Review Section Oxidation Reduction Reactions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review Section Oxidation Reduction Reactions Answers
 - Personalized Recommendations
 - Concept Review Section Oxidation Reduction Reactions Answers User Reviews and Ratings

- Concept Review Section Oxidation Reduction Reactions Answers and Bestseller Lists
- 5. Accessing Concept Review Section Oxidation Reduction Reactions Answers Free and Paid eBooks
 - Concept Review Section Oxidation Reduction Reactions Answers Public Domain eBooks
 - Concept Review Section Oxidation Reduction Reactions Answers eBook Subscription Services
 - Concept Review Section Oxidation Reduction Reactions Answers Budget-Friendly Options
- 6. Navigating Concept Review Section Oxidation Reduction Reactions Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Concept Review Section Oxidation Reduction Reactions Answers Compatibility with Devices
 - Concept Review Section Oxidation Reduction Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Concept Review Section Oxidation Reduction Reactions Answers
 - Highlighting and Note-Taking Concept Review Section Oxidation Reduction Reactions Answers
 - Interactive Elements Concept Review Section Oxidation Reduction Reactions Answers
- 8. Staying Engaged with Concept Review Section Oxidation Reduction Reactions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Concept Review Section Oxidation Reduction Reactions Answers
- 9. Balancing eBooks and Physical Books Concept Review Section Oxidation Reduction Reactions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Concept Review Section Oxidation Reduction Reactions Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Concept Review Section Oxidation Reduction Reactions Answers
 - Setting Reading Goals Concept Review Section Oxidation Reduction Reactions Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concept Review Section Oxidation Reduction Reactions Answers
 - Fact-Checking eBook Content of Concept Review Section Oxidation Reduction Reactions Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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