

CHEMISTRY: COUNTING ATOMS IN COMPOUNDS WORKSHEET #7.0.1

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: $5\text{P}_2\text{O}_3$ $\text{P} = \underline{(5 \times 2 =) 10}$ $\text{O} = \underline{(5 \times 3 =) 15}$

For example: $4\text{Zn}(\text{NO}_3)_2$ $\text{Zn} = \underline{(4 \times 1 =) 4}$ $\text{N} = \underline{(4 \times 1 \times 2 =) 8}$ $\text{O} = \underline{(4 \times 3 \times 2 =) 24}$

1. $4\text{K}_2\text{CO}_3$ $\text{K} = \underline{8}$ $\text{C} = \underline{4}$ $\text{O} = \underline{12}$
2. $2\text{Sr}_3(\text{PO}_4)_2$ $\text{Sr} = \underline{6}$ $\text{P} = \underline{4}$ $\text{O} = \underline{16}$
3. $3\text{N}_4\text{O}_{10}$ $\text{N} = \underline{12}$ $\text{O} = \underline{30}$
4. $2(\text{NH}_4)_3\text{N}$ $\text{N} = \underline{8}$ $\text{H} = \underline{24}$
5. $8\text{Cl}_2\text{O}$ $\text{Cl} = \underline{16}$ $\text{O} = \underline{8}$
6. $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ $\text{Ca} = \underline{1}$ $\text{C} = \underline{4}$ $\text{H} = \underline{6}$ $\text{O} = \underline{4}$
7. 12NaBr $\text{Na} = \underline{12}$ $\text{Br} = \underline{12}$

Counting Atoms In Compounds Answer Naoh

**Francis Landey Patton, Howard Crosby
Butler, Varnum Lansing Collins**

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