

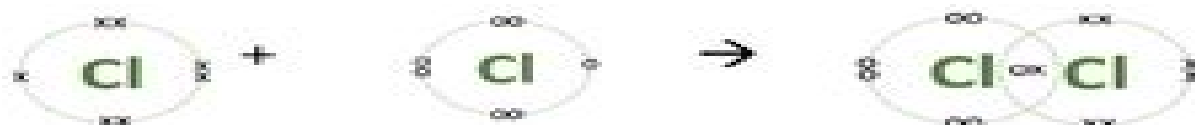
Covalent Bonding

Non-metals need to gain more electrons to have a full outer shell. They can do this in 2 ways:

- **ionic bonding** - atoms lose or gain electrons, forming positive and negative ions which are then electrostatically attracted to each other (an ionic bond). This occurs between a metal and a non-metal.
- **Covalent bonding** - atoms share a pair of electrons (a covalent bond). This occurs between non-metals only.

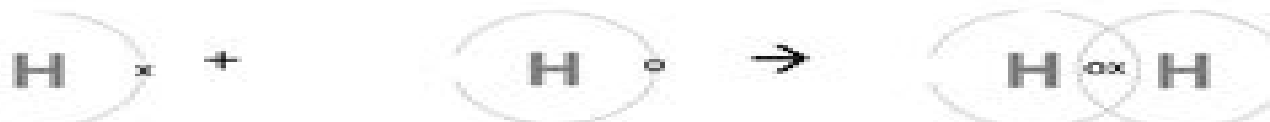
Covalent bonding can only happen between **non-metals** - it does not involve metals.

The 2 atoms "**share**" a pair of electrons. They do this by **overlapping** their outer shells so that the shared electrons are present in the outer shells of **both** atoms. This **shared pair** of electrons is a **covalent bond**.



We can see that as both atoms now possess the shared pair, both atoms have **full outer shells** of **8** electrons in their outer shell (2 shared in the covalent bond, 6 each on their outer shells).

When **hydrogen** atoms are involved in covalent bonding, they will only have **2** electrons in their outer shell as the first shell can only hold 2. However their outer shell is still **full**.



How to draw covalent dot/cross diagrams:

1) How many electrons do the atoms need to gain to have a full outer shell? This is how many bonds (x/c) go in the centre.

E.g. Fluorine is in group 7, so it has 7 electrons in its outer shell. It needs to gain **1** electron so it will form **1** covalent bond. This means we need to draw **1** bond (x/c) in the centre.



2) Now add more electrons to the outer shell of each atom, so that they both have a **full outer shell** (in this case, 8 electrons). Remember that **both** atoms already "**own**" the 2 shared electrons in the centre.

E.g. Each atom already has 2 electrons (from the shared pair), so they each need **6** more. Their outer shells are now **full**.



Covalent Bonding Core Teaching Resources Test Answer

Wenbin Ji



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