

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the hydrogen bond is a weaker attraction. Hydrogen bonds are generally intermolecular, while ionic and covalent bonds occur between ions or atoms respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond. Increased polarity in H₂O bonds means a stronger intermolecular attraction, making water a liquid at room temperature. Hydrogen bonding exists between water molecules, but not between hydrogen sulfide molecules.

Covalent Bonding Study Guide Content Mastery Answers

Henry Sinclair Pickering



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Glencoe Science McGraw-Hill Staff, 2001-06 *Biology* Glencoe/McGraw-Hill, 1994-07 *Introduction to Organic Chemistry* William H. Brown, Thomas Poon, 2016-01-13 Introduction to Organic Chemistry 6th Edition provides an introduction to organic chemistry for students who require the fundamentals of organic chemistry as a requirement for their major. It is most suited for a one semester organic chemistry course. In an attempt to highlight the relevance of the material to students, the authors place a strong emphasis on showing the interrelationship between organic chemistry and other areas of science, particularly the biological and health sciences. The text illustrates the use of organic chemistry as a tool in these sciences; it also stresses the organic compounds, both natural and synthetic, that surround us in everyday life in pharmaceuticals, plastics, fibers, agrochemicals, surface coatings, toiletry preparations, and cosmetics, food additives, adhesives, and elastomers. This text is an unbound, three-hole punched version. Access to WileyPLUS sold separately. **Brown's Introduction to Organic Chemistry** William H. Brown, Thomas Poon, 2017-06-28 Introduction to Organic Chemistry 6th Global Edition provides an introduction to organic chemistry for students who require the fundamentals of organic chemistry as a requirement for their major. It is most suited for a one semester organic chemistry course. In an attempt to highlight the relevance of the material to students, the authors place a strong emphasis on showing the interrelationship between organic chemistry and other areas of science, particularly the biological and health sciences. The text illustrates the use of organic chemistry as a tool in these sciences; it also stresses the organic compounds, both natural and synthetic, that surround us in everyday life in pharmaceuticals, plastics, fibers, agrochemicals, surface coatings, toiletry preparations, and cosmetics, food additives, adhesives, and elastomers. **Handbook of Special Education** James M. Kauffman, Daniel P. Hallahan, 2011-05-15 Special education is now an established part of public education in the United States by law and by custom. However, it is still widely misunderstood and continues to be dogged by controversies related to such things as categorization, grouping, assessment, placement, funding, instruction, and a variety of legal issues. The purpose of this 13-part, 57-chapter handbook is to help profile and bring greater clarity to this sprawling and growing field. To ensure consistency across the volume, chapter authors review and integrate existing research, identify strengths and weaknesses, note gaps in the literature, and discuss implications for practice and future research. Key features include Comprehensive Coverage: Fifty-seven chapters cover all aspects of special education in the United States, including cultural and international comparisons. Issues/Trends: In addition to synthesizing empirical findings and providing a critical analysis of the status and direction of current research, chapter authors discuss issues related to practice and reflect on trends in thinking. Categorical Chapters: In order to provide a comprehensive and comparative treatment of the twelve categorical chapters in section IV, chapter authors were asked to follow a consistent outline: Definition, Causal Factors, Identification, Behavioral Characteristics, Assessment, Educational Programming, and Trends and Issues. Expertise: Edited by two of the most accomplished scholars in special education, chapter

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