

# Student Exploration: Covalent Bonds


**Vocabulary:** covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron


**Prior Knowledge Questions** (Do these BEFORE using the Gizmo.)

1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker. What can they do so that each has a full set of markers? **Frank should give the red marker to Flora, and Flora should give the blue marker to Frank.**
2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set? **Otto should give the black and brown markers to Olivia, and Olivia should give the purple and green markers to Otto.**

## Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo™, atoms form bonds in this way.

To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** (  ) to see the electrons orbiting the nucleus of each atom.

1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? **7 valence electrons.**
2. Click **Pause** (  ). Drag an electron from the left atom to the right atom. Click **Play**.  
What happens? **They equally share the one atom on the new shell that is formed.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.  
What happens now? **They continue to spiral around both atoms, but now there are two electrons on the shell.**



**Activity A:**  
**Sharing electrons**

Get the Gizmo ready:

- Click **Reset**.
- Select **Hydrogen**.



# Covalent Bonds Gizmo

**Guillaume Favre**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

## Covalent Bonds Gizmo:

**The Covalent Bond** Henry Sinclair Pickering, 1977      **New Scientist**, 2008      **Chemical Bonds and Bonds Energy** R Sanderson, 2012-12-02 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists      **Covalent Bonds | Characteristics of Covalent Bonds and Properties of Covalent Compounds | Grade 6-8 Physical Science** Dot EDU, 2024-04-15 Dive into the world of chemistry with this essential guide which is perfect for middle schoolers It unravels the complexities of covalent bonds where atoms share electrons to create molecules and how these interactions form diverse substances Ideal for educators homeschooling parents and school librarians this book emphasizes the significance of understanding chemical bonds within the US STEM curriculum Explore covalent compounds fascinating properties and characteristics through engaging explanations and examples This opportunity will enrich your science lessons and encourage a more profound interest in Chemistry      Chemical Bonding and the Geometry of Molecules George E. Ryschkewitsch, 1963      **Covalent Bond 35 Success Secrets - 35 Most Asked Questions on Covalent Bond - What You Need to Know** Shirley Soto, 2014-10-14 A Blue Ribbon Covalent bond Guide A covalent bond is a biochemical bond that includes the parting of negatron matches amid particles The steady level of alluring and hideous drives amid particles once they share electrons is recognized like covalent joining For numerous particles the parting of electrons permits every one particle to attain the equal of a complete outside shell comparable to a steady microelectronic arrangement There has never been a Covalent bond Guide like this It contains 35 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Covalent bond A quick look inside of some of the subjects covered Chemical bonding Covalent bond Noncovalent bonding Drug Design Covalent bond History Noncovalent bonding Cation Anion Noncovalent bonding Hydrophobic effect Polar covalent bond Polarity of bonds Noncovalent bonding Boiling Points of Liquids Noncovalent bonding London Dispersion Forces Coordinate Covalent Bond Examples Noncovalent bonding effects Noncovalent bonding H

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electronic and crystal structure crystallinity and other states of matter The concluding chapters are devoted to some analytical techniques for structure determination such as diffraction and spectroscopic methods This book is of value to high school and college chemistry teachers and students

**Covalent Bonding in Crystals, Molecules, and Polymers** James C. Phillips, 1969

*The Study and Use of Covalent Bonds in Chemistry: New Insights and Old Knowledge* Ramesh Chandra, Snigdha Singh, 2023-12 Covalent bonds are a fundamental concept in chemistry They represent strong connections between atoms formed by sharing electron pairs These bonds are crucial in shaping molecules structures properties and reactivity Through covalent bonding atoms achieve stable electron configurations resulting in various compounds and materials with unique traits Studying covalent bonds offers valuable insights into chemical bonding intricacies It helps scientists design new molecules comprehend chemical reactions and develop materials for different uses *The Study and Use of Covalent Bonds in Chemistry New Insights and Old Knowledge* is a comprehensive book that explores covalent bonding It provides historical context foundational principles and current research offering a deep understanding of these fundamental interactions By explaining molecular structure bond energy and bond polarity intricacies it helps readers grasp covalent bonds remarkable properties Covering topics like molecular orbital theory intermolecular forces and chemical reactions the book demonstrates covalent bonds versatility and practical applications in various scientific areas

**Let's Bond Together! Explaining Why Atoms Bond, Types of Bonding and Electron Dot Diagrams | Grade 6-8 Physical Science** Dot EDU, 2024-04-15 Dive into the captivating world of chemical bonds with this essential guide for grades 6-8 a must have for the US STEM curriculum This book demystifies the intricate dance of atoms as they bond to form the myriad substances that make up our universe From the basics of covalent and ionic bonds to the detailed exploration of electron dot diagrams it's an invaluable resource for teachers homeschooling parents and librarians Discover the building blocks of matter and how they come together in complex ways Add this to your library to spark a love of science in young learners

Chemical Bonding M.S. Sethi & M. Satake, 2010 Contents Chemical Bonding I Basic Concepts Chemical Bonding II Additional Aspects Intermolecular Force and Crystal Structures

**The Hydrogen Bond** Aloys Hüttermann, 2019-04-01 The author illustrates why the rather weak hydrogen bond is so essential for our everyday life in a lively and entertaining way The chemical and physical fundamentals are explained with examples ranging from the nature of water over the secret of DNA to adhesives and modern detergents The interdisciplinary science is easy to understand and hence a great introduction for chemists biologists and physicists

Acids, Bases, and the Chemistry of the Covalent Bond Calvin Anthony VanderWerf, 1961

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