



CHEMICAL BONDS



IONIC, COVALENT, & METALLIC

FORMS OF CHEMICAL BONDS
Bonding: combination of atoms

Atoms combine to form molecules or compounds by sharing or transferring electrons. This process is called bonding. The resulting molecule or compound has properties different from the individual atoms.

Types of Bonds:

- Covalent Bond:** Formed by sharing electrons between two atoms. Example: H_2 , H_2O , CO_2 .
- Ionic Bond:** Formed by the transfer of electrons from one atom to another, creating oppositely charged ions that attract each other. Example: $NaCl$, MgO .
- Metallic Bond:** Formed by the attraction between positive metal ions and a sea of delocalized electrons. Example: Fe , Cu , Al .

Covalent, Ionic, and Metallic Bonding

Covalent bonding occurs when two atoms share one or more pairs of electrons. This results in a stable molecule. Ionic bonding occurs when one atom transfers an electron to another atom, creating a positive ion and a negative ion. These ions are attracted to each other. Metallic bonding occurs in metals, where the atoms are arranged in a regular lattice, and the outer electrons are delocalized, forming a 'sea of electrons' that holds the metal together.

Science Real

A simplified periodic table of elements, color-coded by groups.

Diagram of Ionic Bonding: Shows a red circle with a minus sign ($-$) and a yellow circle with a plus sign ($+$) connected by a dashed line, representing an ionic bond.

Diagram of Covalent Bonding: Shows two atoms, each with a nucleus and a shell of electrons. A red arrow points to the shared electron pair between the two atoms, representing a covalent bond.

Diagram of Metallic Bonding: Shows a regular lattice of positive metal ions (represented by blue circles) surrounded by a sea of delocalized electrons (represented by a blue cloud).

Text on the right:

...forming a bond in a diatomic molecule. If each atom shares an electron with the other, both atoms have eight valence electrons.

...single, double, or triple covalent bonds. Triple and double bonds are stronger and shorter.

GUIDED READING ACTIVITY

Covalent Bonds Reading Notetaking Guide

Herbert Drapkin



Covalent Bonds Reading Notetaking Guide:

Reading and Writing in Science Maria C. Grant, Douglas Fisher, 2010 Written by a science educator and a literacy expert this resource gives secondary science teachers an approach for developing students disciplinary literacy so they can access science content

Reader's Guide to the History of Science Arne Hessenbruch, 2013-12-16 The Reader's Guide to the History of Science looks at the literature of science in some 550 entries on individuals Einstein institutions and disciplines Mathematics general themes Romantic Science and central concepts Paradigm and Fact The history of science is construed widely to include the history of medicine and technology as is reflected in the range of disciplines from which the international team of 200 contributors are drawn

Focus on Physical Science California Edition Michael J. Padilla, 2007

Chemical Interactions, 2005

Reader's Guide to Periodical Literature Supplement, 1954

Instructor's Guide for Campbell's Biology Nina Caris, Harold T. Underwood, 1996

Drexel Polymer Notes, 1990

Manual of the Public Examinations Board University of Adelaide. Public Examinations Board, 1967 The Manuals include information on syllabus regulations copies of examination papers and notes by examiners They also include pass lists

Student Self-study Guide Donald W. Shive, 1975

Films and Other Materials for Projection Library of

Congress, 1963

Library of Congress Catalog: Motion Pictures and Filmstrips

Library of Congress, 1963 A cumulative list of works represented by Library of Congress printed cards

Readers' Guide to Periodical Literature Anna Lorraine

Guthrie, Bertha Tannehill, Neltje Marie Tannehill Shimer, 1988 An author subject index to selected general interest periodicals of reference value in libraries

Teacher's Manual and Resource Guide for Exploring the Sciences Herbert Drapkin, 1964

Introduction to Modern Inorganic Chemistry Kenneth Malcolm Mackay, Rosemary Ann Mackay, 1981 The electronic structure and the properties of atoms Covalent molecules diatomics Polyatomic covalent molecules The solid state Solution chemistry Experimental methods General properties of the elements in relation to the periodic table Hydrogen The s elements The scandium group and the lanthanides The actinide elements The transition metals general properties and complexes The transition elements of the first series The elements of the second and third transition series Transition metals selected topics The elements of the p block

Holt Chemistry Ralph Thomas Myers, 2004

Science Model Curriculum Guide, Kindergarten Through Grade Eight California. State Department of Education, 1987 This guide was developed with the intention of helping teachers and school site administrators in California review the elementary science curriculum and compare it to an idealized model that is presented in the document Part I of the guide provides a summary of a number of characteristics considered to be important to a strong elementary science program It was designed to aid teachers principals and parents in identifying features of their local science program where attention is needed Part II presents a full scale portrait of an elementary science program that focuses on the development of student understanding This section presents teaching ideas that are concerned with both the knowledge base and science process skills Common themes are present in

the discussion of science instruction in the various subject areas The disciplines and associated themes addressed are 1 biological science cells genetics evolution plants protists animals human beings ecosystems 2 earth science astronomy geology and natural resources meteorology oceanography and hydrology and 3 physical science matter mechanics energy sources and transformation heat light electricity magnetism and sound TW *Modules* McDougal Littell Incorporated,2005

The Macmillan Guide to Correspondence Study Modoc Press,1996 *How to Study Science* Fred Drewes,2000 This text aims to help students get the most out of their science course by giving them suggestions on notetaking managing study time and taking tests A multidisciplinary approach is taken including examples from biology chemistry physics geology and meteorology *Paperbound Book Guide for Colleges* ,1970

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