

IONIC AND COVALENT BOND SCAVENGER HUNT!

Ionic vs. Covalent

Understanding Ionic and Covalent Bonding

Ionic bonds with each other to become more stable. Most atoms want to have eight electrons in their outer energy level (called the outer shell), and bonding helps them achieve this goal. There are two main types of chemical bonding: ionic bonding and covalent bonding. Each type of bond forms in a different way and involves different types of elements.

Ionic Bonding

Ionic bonds happen between a metal and a nonmetal. In this type of bond, electrons are transferred from one atom to another. The metal atom loses electrons and becomes a positive ion (called a cation), while the nonmetal atom gains electrons and becomes a negative ion (called an anion). These oppositely charged ions are attracted to each other, forming a strong ionic bond.

To write the formula of an ionic compound, we use the valuations numbers (charges) of each element. For example, sodium (Na) has a +1 charge and chlorine (Cl) has a -1 charge, so they combine to form NaCl. If the charges are not equal, we use the crisscross method.

For example, magnesium (Mg) is +2 and oxygen (O) is -2. They will form a 1:1 ratio. For example, magnesium (Mg) is +2 and oxygen (O) is -2, we crisscross the numbers (Mg₂O₂).



For aluminum (Al) is +3 and oxygen is -2, we crisscross the numbers (Al₂O₃ subscript added). The 1 goes to 0 and the 2 goes to Al.

Writing ionic compounds is simple: name the metal first, and then the nonmetal, changing the end of the nonmetal to -ide. So NaCl is "sodium chloride" and MgO is "magnesium oxide." If a metal has more than one possible charge (like iron), we show the charge with Roman numerals, like FeCl₂ = iron(II) chloride.

Ionic vs. Covalent

Covalent Bonding

Covalent bonds happen between two nonmetals. Instead of transferring electrons, these atoms share electrons to reach stability. Each shared pair of electrons creates one covalent bond. Some elements share one, two pairs (called a double bond), or even three pairs of electrons (called a triple bond).

Writing formulas for covalent compounds, we use the system to tell us how many atoms of each element are present. That's where Greek prefixes come in.

Prefixes

1 - mono	6 - hexa
2 - di	7 - hepta
3 - tri	8 - octa
4 - tetra	9 - nona
5 - penta	10 - deca

For example, 2 oxygen, 7 hydrogen, 1 nitrogen, 1 carbon is N₂H₇O₂C.

Remember, the "1" at the end of the word is left off. Example: 1 oxygen, 2 hydrogen, 1 nitrogen, 1 carbon is N₂H₂O₂C.

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