## Chem 105: Exam 3 Study Guide

#### October 19th, 2020 Lecture

#### **Basics of Quantum Mechanics:**

Quantum Mechanics -- focuses on electrons or photons

#### Litolat

- is electromagnetic radiation
- is wave-like and particle-like
- is not continuous with infinite values, it can only have specific values (shown through visible spectrum)
- must overcome binding energy of an electron to eject it from a metal
- only has very specific values

Frequency threshold has to be met to achieve the photoelectric effect

#### Electromagnetic Spectrum:

Visible light

- Violet → lowest visible wavelength
- Red → highest visible wavelength
- Lowest energy ← → highest energy

Each element gives off electrons at specific light energies

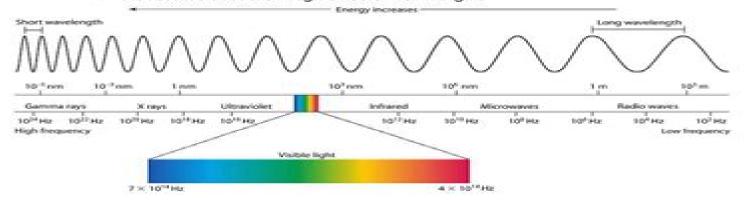
- can differentiate elements based on this
- light is not continuous because of this

The further out the electron is the higher energy it gives off

- High end → lower one : emit light
- Low level → higher one : absorbs light

Only a few light colors show up in visible light when reading colors for elements

- ROYGBIV
- the rest are shown through all other wavelengths



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#### **Table of Contents Chem 105 Study Guide**

- 1. Understanding the eBook Chem 105 Study Guide
  - ∘ The Rise of Digital Reading Chem 105 Study Guide
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chem 105 Study Guide
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chem 105 Study Guide
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chem 105 Study Guide
  - Personalized Recommendations
  - Chem 105 Study Guide User Reviews and Ratings
  - Chem 105 Study Guide and Bestseller Lists
- 5. Accessing Chem 105 Study Guide Free and Paid eBooks
  - Chem 105 Study Guide Public Domain eBooks

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- Chem 105 Study Guide Budget-Friendly Options
- 6. Navigating Chem 105 Study Guide eBook Formats
  - o ePub, PDF, MOBI, and More
  - Chem 105 Study Guide Compatibility with Devices
  - Chem 105 Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chem 105 Study Guide
  - Highlighting and Note-Taking Chem 105 Study Guide
  - Interactive Elements Chem 105 Study Guide
- 8. Staying Engaged with Chem 105 Study Guide
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chem 105 Study Guide
- 9. Balancing eBooks and Physical Books Chem 105 Study Guide
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chem 105 Study Guide
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chem 105 Study Guide
  - $\circ\,$  Setting Reading Goals Chem 105 Study Guide
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chem 105 Study Guide
  - Fact-Checking eBook Content of Chem 105 Study Guide
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

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- Interactive and Gamified eBooks

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