

# CHAPTER 14

## CHEMICAL EQUILIBRIUM

### Problem Categories

*Biological:* 14.102, 14.128.

*Conceptual:* 14.13, 14.23, 14.39, 14.53, 14.54, 14.55, 14.56, 14.57, 14.58, 14.59, 14.60, 14.61, 14.62, 14.66, 14.68, 14.69, 14.70, 14.71, 14.83, 14.93, 14.99, 14.100, 14.110, 14.111, 14.124.

*Descriptive:* 14.87, 14.92, 14.95.

*Environmental:* 14.67, 14.81.

*Industrial:* 14.81, 14.97, 14.101, 14.105.

### Difficulty Level

*Easy:* 14.14, 14.15, 14.16, 14.17, 14.18, 14.19, 14.20, 14.21, 14.22, 14.29, 14.30, 14.36, 14.53, 14.54, 14.55, 14.56, 14.58, 14.59, 14.60, 14.63, 14.69, 14.79, 14.96, 14.102, 14.103, 14.107.

*Medium:* 14.13, 14.23, 14.25, 14.26, 14.28, 14.31, 14.32, 14.35, 14.39, 14.40, 14.41, 14.42, 14.43, 14.44, 14.45, 14.48, 14.57, 14.61, 14.62, 14.64, 14.65, 14.66, 14.67, 14.68, 14.70, 14.71, 14.72, 14.76, 14.77, 14.80, 14.83, 14.84, 14.86, 14.87, 14.89, 14.92, 14.94, 14.95, 14.97, 14.99, 14.100, 14.101, 14.105, 14.111, 14.112, 14.118, 14.121, 14.125.

*Difficult:* 14.24, 14.27, 14.46, 14.47, 14.73, 14.74, 14.75, 14.78, 14.81, 14.82, 14.85, 14.88, 14.90, 14.91, 14.93, 14.98, 14.104, 14.106, 14.108, 14.109, 14.110, 14.113, 14.114, 14.115, 14.116, 14.117, 14.119, 14.120, 14.122, 14.123, 14.124.

**14.13**      $K_c = \frac{[B]}{[A]}$

- (1) With  $K_c = 10$ , products are favored at equilibrium. Because the coefficients for both A and B are one, we expect the concentration of B to be 10 times that of A at equilibrium. Choice (a) is the best choice with 10 B molecules and 1 A molecule.
- (2) With  $K_c = 0.10$ , reactants are favored at equilibrium. Because the coefficients for both A and B are one, we expect the concentration of A to be 10 times that of B at equilibrium. Choice (d) is the best choice with 10 A molecules and 1 B molecule.

You can calculate  $K_c$  in each case without knowing the volume of the container because the mole ratio between A and B is the same. Volume will cancel from the  $K_c$  expression. Only moles of each component are needed to calculate  $K_c$ .

**14.14** Note that we are comparing similar reactions at equilibrium – two reactants producing one product, all with coefficients of one in the balanced equation.

- (a) The reaction,  $A + C \rightleftharpoons AC$  has the **largest** equilibrium constant. Of the three diagrams, there is the most product present at equilibrium.
- (b) The reaction,  $A + D \rightleftharpoons AD$  has the **smallest** equilibrium constant. Of the three diagrams, there is the least amount of product present at equilibrium.

**14.15** When the equation for a reversible reaction is written in the opposite direction, the equilibrium constant becomes the reciprocal of the original equilibrium constant.

$$K' = \frac{1}{K} = \frac{1}{4.17 \times 10^{-34}} = 2.40 \times 10^{33}$$

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**David Baud**



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