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A properly adjusted Bunsen flame is a blue, nonluminous flame.

Experiment 1

Basic Laboratory Operations

- To light and properly adjust the flame of a Bunsen burner
- To develop the skill for properly operating a balance
- To develop the technique of using a pipet
- To determine the density of an unknown substance

OBJECTIVES

The following techniques are used in the Experimental Procedure:



TECHNIQUES

You will use a number of techniques repeatedly throughout your laboratory experience. Seventeen principal techniques are fully described under **Laboratory Techniques** in this manual (pp. 17–42). The application of a technique in the Experimental Procedure is noted in the page margin by the corresponding technique icon. Become familiar with each technique and icon before using it in the experiment.

In this experiment, you will learn several common techniques that are used repeatedly throughout this laboratory manual. You will learn to light and adjust a Bunsen burner, to use a laboratory balance, and to use a pipet. With the skills developed in using a balance and pipet, you will determine the density of a metal and a liquid.

INTRODUCTION

Laboratory burners come in many shapes and sizes, but all accomplish one main purpose: A combustible gas–air mixture yields a hot, efficient flame (Figure 1.1, page 52). The combustible gas used to supply the fuel for the Bunsen burner in most laboratories is natural gas. Natural gas is a mixture of gaseous hydrocarbons, but primarily the hydrocarbon methane, CH_4 . If sufficient oxygen is supplied, methane burns with a blue, **nonluminous** flame, producing carbon dioxide and water as combustion products:



With an insufficient supply of oxygen, small carbon particles are produced, which, when heated to **incandescence**, produce a yellow, luminous flame. The combustion products may, in addition to carbon dioxide and water, include carbon monoxide.

Bunsen Burner

Robert Bunsen, 1811–1899

Hydrocarbon: a molecule consisting of only the elements carbon and hydrogen

Nonluminous: nonglowing or nonluminescent

Incandescence: glowing with intense heat

Chem 11lab Manual Answers Experiment 1

Ying Liu



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Laboratory Manual for Principles of General Chemistry Jo Allan Beran, 2010-11-01 This new edition of the Beran lab manual emphasizes chemical principles as well as techniques The manual helps students understand the timing and situations for the various techniques The Beran lab manual has long been a market leading lab manual for general chemistry Each experiment is presented with concise objectives a comprehensive list of techniques and detailed lab intros and step by step procedures

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Guided Inquiry Experiments for General Chemistry Nancy K. Kerner, Ram S. Lamba, 2007-10-19 The use of the laboratory is a valuable tool in developing a deeper understanding of key chemical concepts from the experimental process This lab manual encourages scientific thinking enabling readers to conduct investigations in chemistry It shows how to think about the processes they are investigating rather than simply performing a laboratory experiment to the specifications set by the manual Each experiment begins with a problem scenario and ends with questions requiring feedback on the problem

Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-09-22 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

The Journal of Industrial and Engineering Chemistry, 1921 *Lab Manual to Accompany Introduction to Chemistry* William L. Masterton, Stanley M. Cherim, 1984-02

Journal of Industrial and Engineering Chemistry, 1921 Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition

focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order

Laboratory Manual to Accompany Chemistry in Context American Chemical Society,2005-02 The 5th edition Laboratory Manual that accompanies Chemistry in Context is compiled and edited by Gail Steehler Roanoke College The experiments use microscale equipment wellplates and Beral type pipets as well as common materials Project type and cooperative collaborative laboratory experiments are included Additional experiments are available on the Online Learning Center as is the instructor s guide **Journal of Chemical Education** ,1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers **Introduction to Chemistry, Laboratory Manual** T. R. Dickson,1994-12-23 Teaches chemistry by offering a dynamic provocative and relevant view of the topic and its importance to society and our daily lives Three themes are stressed throughout the text developing chemical thinking and a chemical vision learning problem solving methods and utilizing group work and discussion activities These themes involve and engage the students in their own learning processes they are challenged to be active The presentation of topics has been altered to include a new chapter which introduces the students to scientific thinking and shows that chemistry involves interesting and relevant topics The reorganization presents many core concepts in the first five chapters preparing students for later chapters In addition the author has added vignettes throughout the chapters referring to health technology the environment and society as well as to specific tools of direct use to students

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