

Part I: Potential Energy Diagrams

Read pages 527-529 in *Chemistry: the Central Science*.

The energy changes involved in chemical reactions can be conveniently examined using potential energy diagrams. The total potential energy of all the reactants in the reaction is illustrated at the left of the diagram, the total potential energy of the products is at the right side, and the activation energy change or changes are illustrated in between.

On the grids below, plot energy diagrams illustrating the reactions described at left.

Reaction 1:

Potential energy of reactants = 250 kJ

Potential energy of activated complex = 350 kJ

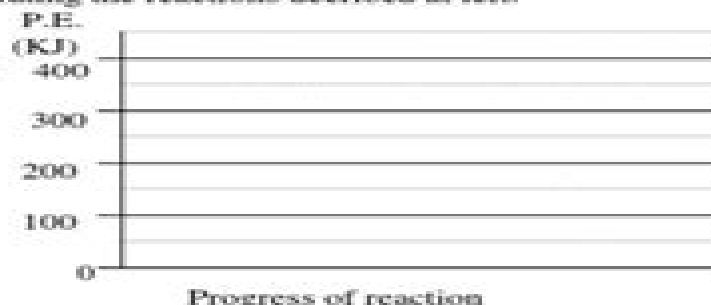
Potential energy of products = 300 kJ

Is this reaction exothermic or endothermic?

How can you tell?

What is the value of ΔH for this reaction?

What is the activation energy for this reaction?



In the presence of a catalyst for this reaction, how would the diagram change? Explain which of the following would be changed, and how: energy of reactants, energy of products, energy of activated complex, rate of reaction.

Reaction 2:

Potential energy of reactants = 350 kJ

Activation energy = 100 kJ

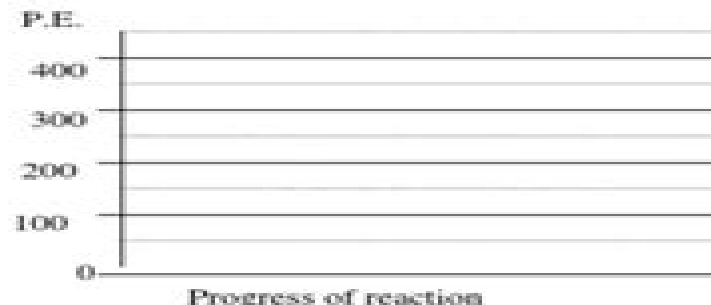
Potential energy of products = 250 kJ

Is this reaction exothermic or endothermic?

How can you tell?

What is the value of ΔH for this reaction?

What is the energy of the activated complex for this reaction?



Suppose the concentrations of reactants were increased. Explain which of the following would be changed, and how: energy of reactants, energy of products, energy of activated complex, rate of reaction.

Selected Answers: Rxn 1: Endothermic... products at higher PE than reactants. $\Delta H = +50$ kJ, $E_a = 100$ kJ. With catalyst, E_a would be less, increasing the rate of reaction. Rxn 2: exothermic... products at lower PE. $\Delta H = -100$ kJ. PE of activated complex = 450 kJ. Conc. Changes would affect the rate only. Rxn 4 (next page): PE of reactants = 100 kJ, of activated complex = 350 kJ, of products = 50 kJ. $\Delta H = -50$ kJ, $E_a = 250$ kJ

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