

Meet and Greet: Exchange email and/or cell number with at least 2 classmates

$$S = k_B(P_A - P_{\text{total}}) = X_A(P_A) + X_B(P_B), \quad \Delta T = K_{\text{eb}}(m), \quad q = m(C_p)\Delta T.$$

1.) Which one of the following has a low density?

- A) gas B) liquid C) solid D) none of the above E) all of the above

2.) Place the following compounds in order of decreasing strength of intermolecular forces.

- I. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$ II. $(\text{CH}_3)_3\text{CCH}_3$ III. $(\text{CH}_3)_2\text{CCH}_2\text{CH}_3$
 A) III > II > I B) I > III > II C) I > II > III D) II > III > I E) III > I > II

The more surface area, the stronger the London dispersion forces. Chain-like structures will have much higher surface area compared to branch-like and cyclical structures of similar molar masses.

3.) Identify the compound with the highest dipole moment.

- A) CH_3OCH_3 B) $\text{CH}_3\text{CH}_2\text{CH}_3$ C) CH_3CHO D) $(\text{CH}_3\text{CH}_2)_2\text{O}$ E) CH_3CN

CH_3CN has a linear 3-D environment separating its polar and non-polar regions, giving it a large dipole moment. CH_3CHO has a trigonal planar 3-D environment separating its polar and non-polar regions, making it have a slightly lower net dipole moment. It is true that a C-O bond is more polar than a C-N bond, but the overall shape of the molecule plays a large role in determining the net dipole moment.

4.) Give the major force between ethanol (CH_3OH) and water.

- A) dipole-dipole
 B) dispersion
C) hydrogen bonding definition of hydrogen bonding is a X-H bond (X = O, N, F)
 D) ion-ion
 E) ion-dipole

Remember that for intermolecular forces hydrogen bonding > dipole-dipole > London/dispersion

5.) Give the major force between acetone ($\text{CH}_3\text{C}(\text{O})\text{CH}_3$) and chloroform (CHCl_3).

- A) dipole-dipole Both have a dipole
 B) dispersion
 C) hydrogen bonding
 D) ion-ion
 E) ion-dipole

Remember that for intermolecular forces hydrogen bonding > dipole-dipole > London/dispersion

6.) Define *critical point*.

- A) The temperature and pressure above which a supercritical fluid exists.
 B) The temperature and pressure where liquid, solid, and gas are equally stable and are in equilibrium.
 C) The temperature and pressure below which a supercritical fluid exists.

Chem 162 Lab Manual Answers

Mary Burnham, Carol Hurd



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