

CHEM 104-S01/S02 - Extra Practice: 4 Answers

1. Specify what ions are present upon dissolving each of the following substances in water:



2. Acetone, CH_3COCH_3 , is a non-electrolyte, hypochlorous acid, HClO , is a weak electrolyte, and ammonium chloride, NH_4Cl , is a strong electrolyte.

i) What are the solute particles present in aqueous solutions of each compound?

acetone (a nonelectrolyte): $\text{CH}_3\text{COCH}_3(\text{aq})$ molecules only; hypochlorous acid (a weak electrolyte): $\text{HClO}(\text{aq})$ molecules, $\text{H}^{+}(\text{aq})$, $\text{ClO}^{-}(\text{aq})$; ammonium chloride (a strong electrolyte): $\text{NH}_4^{+}(\text{aq})$, $\text{Cl}^{-}(\text{aq})$

ii) If 0.1 mol of each compound is dissolved in solution, which one may theoretically contain 0.2 mol of solute particles, which contains 0.1 mol of solute particles, and which contains somewhere between 0.1 and 0.2 mol of solute particles?

NH_4Cl , 0.2 mol solute particles; HClO , between 0.1 and 0.2 mol solute particles; CH_3COCH_3 , 0.1 mol solute particles

3. Predict which of the following compounds are soluble in water:

i) $\text{Ni}(\text{OH})_2$ **insoluble**

ii) PbBr_2 **insoluble**

iii) $\text{Ba}(\text{NO}_3)_2$ **soluble**

iv) AlPO_4 **insoluble**

v) $\text{AgC}_2\text{H}_3\text{O}_2$ **soluble**

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Table of Contents Chem Activity 4answers

1. Understanding the eBook Chem Activity 4answers
 - The Rise of Digital Reading Chem Activity 4answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Activity 4answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem Activity 4answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Activity 4answers
 - Personalized Recommendations
 - Chem Activity 4answers User Reviews and Ratings
 - Chem Activity 4answers and Bestseller Lists
5. Accessing Chem Activity 4answers Free and Paid eBooks
 - Chem Activity 4answers Public Domain eBooks
 - Chem Activity 4answers eBook Subscription Services
 - Chem Activity 4answers Budget-Friendly Options

6. Navigating Chem Activity 4answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem Activity 4answers Compatibility with Devices
 - Chem Activity 4answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem Activity 4answers
 - Highlighting and Note-Taking Chem Activity 4answers
 - Interactive Elements Chem Activity 4answers
8. Staying Engaged with Chem Activity 4answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Activity 4answers
9. Balancing eBooks and Physical Books Chem Activity 4answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Activity 4answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Activity 4answers
 - Setting Reading Goals Chem Activity 4answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Activity 4answers
 - Fact-Checking eBook Content of Chem Activity 4answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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