

Chemical Reactions and Balancing Chemical Equations Web Quest

Name: _____ Date: _____ Class: _____

Chemical Equations

http://web.archive.org/web/20131203110147/http://www.files.chem.vt.edu/RVGS/ACT/notes/Ty pes_of_Equations.html

Click "Directions"

- What three things does a balanced equation show you?
 - The _____ which enter into a reaction.
 - The _____ which are formed by the reaction.
 - The amounts of each substance _____ and each _____ produced.
- What two things must we remember when balancing equations?
 - Every chemical compound has a _____ which cannot be _____.
 - A chemical _____ must account for every _____ that is used, which is an application of the Law of _____ of _____.
- What does the \longrightarrow mean? _____
- What does the \longleftrightarrow mean? _____

Classic ChemBalancer –

<http://funbasedlearning.com/chemistry/chemBalancer/default.htm>

(1) Click the button for "Directions" and **read carefully**. Click the "OK" button and return to the game screen.

(2) Click "Start Game" button to give it a try!

(3) Start by adding a "1" in each box and compare the number of atoms of each element you have on each side of the equation.

(4) Change coefficients to balance each equation and click the "Balanced" button to check it. If it is wrong, correct it!

(5) Use the information in the pop-up windows to answer each question and then write the balanced equation before clicking the "OK" button.

#1

What does "ferrum" mean? _____
What color is sulfur? _____

#2

What is HCl? _____
Where is it found in the body? _____

#3

What are pyrotechnics? _____

#4

What was the Hindenberg? _____
What gas was used in it? _____
What gas is used today? _____

Chem Quest 28 Chemical Reactions

M Mosston



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Information Sources in Chemistry R. T. Bottle, J. F. B. Rowland, 2011-12-01 The aim of each volume of this series Guides to Information Sources is to reduce the time which needs to be spent on patient searching and to recommend the best starting point and sources most likely to yield the desired information The criteria for selection provide a way into a subject to those new to the field and assists in identifying major new or possibly unexplored sources to those who already have some acquaintance with it The series attempts to achieve evaluation through a careful selection of sources and through the comments provided on those sources *FEDLINK Services Directory for Fiscal Year ...* ,1991 **Comprehensive Medicinal Chemistry** Corwin Hansch, Peter George Sammes, John Bodenhan Taylor, 1990 V 1 General principles volume editor Peter D Kennewell v 2 Enzymes other molecular targets volume editor Peter G Sammes v 3 Membranes receptors volume editor John C Emmett v 4 Quantitative drug design volume editor Christopher A Ramsden v 5 Biopharmaceutics volume editor John B Taylor v 6 Cumulative subject index drug compendium volume editor Colin J Drayton Directory of Online Databases ,1991 Even numbered issues will contain update information odd numbered issues will be complete reissues with all new revised information fully integrated into the basic Directory *Data Base Alert* ,1988 **The Compu-mark Directory of U.S. Trademarks** ,1991 The Directory of U.S. Trademarks ,1993 Bulletin des bibliothèques de France ,1987 Bulletin de documentation bibliographique appears as separately paged section 1959 1964 70 Actions and reactions ,1978 Chemical Reactions Kaufman George Falk, 1920 Atomic Reactions Michael Polanyi, 1932 **Rates of chemical reactions** R. J. Stanley, 1971 **Types of Chemical Reactions** Lifeliqe, 2019 This lesson plan covers general equations for combination decomposition single replacement and double replacement reactions and predicting what kind of reaction will occur

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