

- I.
 - A. For the reaction $3 A \rightarrow B + 2 C$, the rate of change of B is 0.125 M/s . What is the corresponding rate of change of A?
 - B. For the reaction $2 A \rightarrow B + 3 C$, the rate of change of A is $-0.420 \times 10^{-3} \text{ M} \cdot \text{s}^{-1}$. What is the corresponding rate of change of C?
 - C. The reaction $A + 3 B \rightarrow \text{Products}$ has an initial rate of 0.0345 M/s and the rate law rate $= k[A]^2[B]$. What will the initial rate be if [A] is quadrupled and [B] is halved?
2. Determine the rate law given for the following reaction given the initial rate data in the table:



Experiment	[A] (M)	[B] (M)	Initial rate (M/s)
1	0.0422	0.113	2.35×10^{-5}
2	0.0844	0.113	9.40×10^{-5}
3	0.0844	0.226	1.88×10^{-4}

3. The first-order reaction $A \rightarrow B$ has a rate constant of 0.0643 s^{-1} . How many seconds will it take for 35.2% of A to react?
4. The decomposition reaction $AB \rightarrow A + B$ is a second-order process with a rate constant of $6.92 \times 10^{-3} \text{ M}^{-1} \cdot \text{s}^{-1}$. The initial concentration of A is 0.0248 M . Calculate the concentration of A after 78.9 min.

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John W. Hill



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