

1. A solution is made by dissolving some salt in a beaker of water. The salt is referred to as the
A. solute B. filtrate C. solution D. solvent
2. When KCl dissolves in water, the following will be produced.
A. K and Cl B. K^+ and Cl^- C. K and Cl_2 D. K^+ and Cl_2
3. What is the concentration, in % (m/v) of a solution containing 15.0 g KCl in 600.0 mL solution?
A. 5.00% B. 2.60% C. 0.200% D. 2.50%
4. A 5.00 mL sample of solution has 2.8×10^{-4} g of calcium ions. The ppm concentration is
A. 18 ppm B. 56 ppm C. 2.8×10^{-5} ppm D. 2.8×10^7 ppm
5. How many mL of 2.50% (m/v) salt solution would contain 1.30 g of salt?
A. 105 mL B. 104 mL C. 45.0 mL D. 72.0 mL
6. What is the molarity of a solution that contains 1.50 mol HCl in 2.50 L of solution?
A. 1.67 M B. 0.600 M C. 1.20 M D. 1.40 M
7. How many moles of $C_{10}H_{22}O_{11}$ are needed to prepare 2.50 L of 0.300 M solution?
A. 0.750 mol B. 0.430 mol C. 8.33 mol D. 1.20 mol
8. What is the molarity of a $NaNO_3$ solution made by diluting 250.0 mL of a 1.60 M solution to a final volume of 400. mL?
A. 1.20 M B. 1.00 M C. 0.200 M D. 0.160 M
9. In the process known as osmosis, _____ moves through a semi-permeable membrane into an area of lower _____ concentration.
A. water, water B. solute, solute C. water, solute D. solute, water
10. What is the osmolarity of 0.800M $(NH_4)_3PO_4$ solution?
A. 3.20 osmol B. 2.40 osmol C. 1.60 osmol D. 8.00 osmol

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