

CHEM 102 - Exam 1 Study Guide

Week 1

- Describing matter as uniform/nonuniform
- Chemistry is the study of matter
- Describing matter
 - Properties
 - Mass and volume are extensive properties
 - Density and temperature are intensive
 - Chemistry
 - Measurement is based on
 - International System
 - Significant Figures
- Matter's atomic theory (Dalton)
- Early atomic models of tiny indivisible particles called atoms
 - Thomson's atom - atom was indivisible particles
- Discovery of lightest element gas called H₂
 - Thomson's atom - isotopes exist
- Compounds are formed when atoms of different elements combine
 - Law - Law of Definite Proportions
- Chemical reactions and experiments of Dalton's
 - Dalton - what about the electrons?
- Thomson's Discovery of Electrons
 - Cathode ray experiment
 - Discovered the electrons
 - Determined the charge to mass ratio of electrons
 - Cathode ray experiment
 - Determined the charge to mass ratio of electrons
- Atomic Structure
 - Early Models
 - Proposals by J.J. Thomson and Lord Kelvin
 - Plum pudding model - atom is a "soup" of charges
 - Thomson Model
 - But how the negatively charged ions
 - Very positively charged particles throughout the atom
 - Negatively charged ions or negatively charged particles
 - Nature of cathode ray particles
 - Cathode - 1.6×10^{-19} kg
 - Anode - 1.6×10^{-19} kg
 - Cathode - 1.6×10^{-19} kg

Chem 102 Final Exam Unc Study Guide

Chao Zhang



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Table of Contents Chem 102 Final Exam Unc Study Guide

1. Understanding the eBook Chem 102 Final Exam Unc Study Guide
 - The Rise of Digital Reading Chem 102 Final Exam Unc Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 102 Final Exam Unc Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 102 Final Exam Unc Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 102 Final Exam Unc Study Guide
 - Personalized Recommendations
 - Chem 102 Final Exam Unc Study Guide User Reviews and Ratings
 - Chem 102 Final Exam Unc Study Guide and Bestseller Lists

5. Accessing Chem 102 Final Exam Unc Study Guide Free and Paid eBooks
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 - Chem 102 Final Exam Unc Study Guide eBook Subscription Services
 - Chem 102 Final Exam Unc Study Guide Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Chem 102 Final Exam Unc Study Guide Compatibility with Devices
 - Chem 102 Final Exam Unc Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 102 Final Exam Unc Study Guide
 - Highlighting and Note-Taking Chem 102 Final Exam Unc Study Guide
 - Interactive Elements Chem 102 Final Exam Unc Study Guide
8. Staying Engaged with Chem 102 Final Exam Unc Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 102 Final Exam Unc Study Guide
9. Balancing eBooks and Physical Books Chem 102 Final Exam Unc Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 102 Final Exam Unc Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem 102 Final Exam Unc Study Guide
 - Setting Reading Goals Chem 102 Final Exam Unc Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem 102 Final Exam Unc Study Guide
 - Fact-Checking eBook Content of Chem 102 Final Exam Unc Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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