

ACS General Chemistry Cheat Sheet #1

Protons: positively charged atoms

Neutrons: neutral charged atoms

Electrons: negatively charged atoms

Isotope: an atom with the same number of protons, but different number of neutrons.



NOTE: you need to remember the atomic number and atomic mass of an element to calculate certain things.

Questions from the ACS General Chemistry-Second Edition Exam Book

1. An atom of strontium-90, $^{90}_{38}\text{Sr}$ contains,
- a. 38 electrons, 38 protons, and 52 neutrons b. 38 electrons, 38 protons, 90 neutrons
c. 52 electrons, 52 protons, 38 neutrons d. 52 electrons, 38 protons, 38 neutrons

The correct answer for this question would be **a**. The atomic mass number is the sum of protons and neutrons. The atomic number is the amount of neutrons a specific element has. In this case, we have 38 protons which is the same amount for electrons, so we can go ahead and eliminate **choices c** and **d**. If you take the atomic mass number and subtract it from the atomic number, you can get the number of neutrons the element has.

"One thing you do need to know for the exam is how to calculate relative abundance."

2. Chlorine occurs naturally as a mixture of two isotopes with atomic masses of 34.97 amu and 36.97 amu. What are the relative abundances of these isotopes?

Remember: Isotopes are elements with the same number of protons, but different number of neutrons.

The formula that can be used to calculate the answer is $[(\text{mass isotope}_1) \times \frac{\% \text{ isotope 1}}{100\%}] + [(\text{mass isotope}_2) \times \frac{\% \text{ isotope 2}}{100\%}] = \text{average atomic mass}$

$$\frac{35.45 \text{ amu} - 34.97 \text{ amu}}{36.97 \text{ amu} - 34.97 \text{ amu}} \times 100$$

$$100\% - 24\% = 76\%$$

The correct answer is 76%.

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JG Myers



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