

CHEM ACTIVITY

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STEM 12-YA-17

Title: CHEMISTRY SEVENTH LABORATORY EXPERIMENT: THERMOCHEMISTRY

Objectives: In this experiment, NaOH was performed to determine the enthalpy value of chemical reaction and its correlation with the concentration.

Materials:

Deionized water	Volumetric flask	Pipette	Cover
NaOH (0.2 M)	Magnetic stirrer	Digital thermometer	Stopwatch
NaOH (0.3 M)	Magnetic stir band	Yellow Cable	
NaOH (0.5 M)	Digital balance	Tape	

Procedures: First, the weight of NaOH was measured. Then 100 ml of deionized water was taken then was put into the volumetric flask. A beaker glass was taken and it was covered with newspaper twice to keep the temperature of the solution later. Then after the beaker glass was covered up with the newspaper, it was put above the magnetic stirrer. The deionized water was poured into the beaker glass. After that, the digital thermometer was used to measure the temperature. The cable was secured with tape by taping it to the inside of the beaker glass from being directly touched with the base of the beaker glass. Then it was stirred for about 200 rpm and the measurement was started. The temperature was written down every 5 second. When it hit the second of 20, the NaOH was added into the solution. Lastly it was measured until 5 minutes.

Guide Questions: What will happen If solutes absorb energy from the water as they dissolve? What will happen If solutes release energy to the water as they dissolve?

Interpretation: (1) The water gets colder and the reaction is endothermic (2) The water gets warmer and the reaction is exothermic

Results/Conclusion: Based on data, it can be concluded that the ΔH average is 35703.95 J/mol

Chem Activity 53 Answers

American Chemical Society

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Watts' Dictionary of Chemistry Henry Watts, Henry Forster Morley, Matthew Moncrieff Pattison Muir, 1894 Reader's Guide to Periodical Literature Supplement, 1954 **Mastering Chemistry** Gerald W. Gibson, 1975 *Industrial Arts Index*

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