

Exercises

1. Use the vector notation (see Figure 2 above) to designate the dipole moments of:
CO ; HI ; ClF .
2. Write the Lewis structure for NO_3^- . Explain why the dipole moment is zero. (Hint: where does the center of negative charge have to be such that the dipole moment is zero?)
3. Classify each of the following molecules (or ions) as nonpolar (dipole moment = 0) or polar (dipole moment $\neq 0$). (Hint: it is generally a good idea to determine the shape of the molecule.)

O_2 ; I_2 ; CO ; NO ; H_2O ; CO_3^{2-} ; SO_4^{2-} ; SO_2



4. Which molecule of each pair has a dipole moment of zero?

a) N_2 or NO	b) CH_4 or CH_3Cl	c) SO_2 or SO_3
d) NO_3^- or NO_2^-	e) SO_4^{2-} or SO_2	f) CO_2 or OCS
5. Indicate whether the following statement is true or false and explain your reasoning.

The dipole moment of NH_3 is smaller than the dipole moment of CCl_4 .

6. Which molecule of each pair has the greater dipole moment?

- a) CCl_4 or CH_3Cl
- b) CH_3Br or CH_3Cl
- c) NF_3 or NI_3
- d) H_2O or H_2S

7. Based on the data in Table 1, predict the dipole moment of CH_3F .

Chemactivity 23 The Dipole Moment Answers

J Spring



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Table of Contents Chemactivity 23 The Dipole Moment Answers

1. Understanding the eBook Chemactivity 23 The Dipole Moment Answers
 - The Rise of Digital Reading Chemactivity 23 The Dipole Moment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemactivity 23 The Dipole Moment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemactivity 23 The Dipole Moment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemactivity 23 The Dipole Moment Answers
 - Personalized Recommendations
 - Chemactivity 23 The Dipole Moment Answers User Reviews and Ratings
 - Chemactivity 23 The Dipole Moment Answers and Bestseller Lists
5. Accessing Chemactivity 23 The Dipole Moment Answers Free and Paid eBooks
 - Chemactivity 23 The Dipole Moment Answers Public Domain eBooks
 - Chemactivity 23 The Dipole Moment Answers eBook Subscription Services
 - Chemactivity 23 The Dipole Moment Answers Budget-Friendly Options

6. Navigating Chemactivity 23 The Dipole Moment Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemactivity 23 The Dipole Moment Answers Compatibility with Devices
 - Chemactivity 23 The Dipole Moment Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemactivity 23 The Dipole Moment Answers
 - Highlighting and Note-Taking Chemactivity 23 The Dipole Moment Answers
 - Interactive Elements Chemactivity 23 The Dipole Moment Answers
8. Staying Engaged with Chemactivity 23 The Dipole Moment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemactivity 23 The Dipole Moment Answers
9. Balancing eBooks and Physical Books Chemactivity 23 The Dipole Moment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemactivity 23 The Dipole Moment Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemactivity 23 The Dipole Moment Answers
 - Setting Reading Goals Chemactivity 23 The Dipole Moment Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemactivity 23 The Dipole Moment Answers
 - Fact-Checking eBook Content of Chemactivity 23 The Dipole Moment Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

-
- Interactive and Gamified eBooks

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