

chemical analysis

TOPIC 8

PURITY + FORMULATIONS:

Pure: one element/compound throughout

Using boiling point to tell purity:

- 1) Test melting/boiling point of substance
- 2) Compare to melting/boiling point of the pure substance (can be found in a data book) it will be a **sharp point**
- 3) Impurities lower melting point and raise boiling point as well as increase melting/boiling range.

Formulations: mixtures with precise purposes

- each component present in a measured quantity
- used in the pharmaceutical industry to make pills correct part of body, concentration, consumable, and shelf life.
- Also used in fertilisers, cosmetics, fuels, metal alloys. can be expressed as ratio or percentage.

PAPER CHROMATOGRAPHY:

- 1) Equilibrium forms between mobile/stationary phases.
- 2) How quickly a chemical moves is **how it is distributed between the two phases** if it spends more in mobile or stationary phase **more time mobile moves further.**

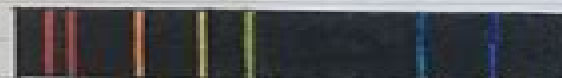
- 3) components in mixture will separate through stationary phase (paper) if they spend different times in the stationary phase
- pure substance** shows in one spot as there's only 1 substance
- R_f** values is the distance traveled by solute **distance traveled by B** / **distance traveled by A**



FLAME EMISSION SPECTROSCOPY:

- 1) Sample placed in flame, electrons heat up and gain energy. When they return to original energy level they transfer energy as light which passes through a **spectroscope**.

- 2) The spectroscope can **detect different wave lengths of light** so each ion produces a different pattern of wave lengths with a different line spectrum. The intensity of the spectrum indicates the concentration of the ion in the solution. Spectra can be used to identify ions and calculate their concentration.



GAS + ANION TESTS:

Gas Tests:

- Chlorine** - Bleaches damp **litmus paper white** (red initially for acids)
- oxygen** - will relight a glowing **spint**
- CO₂** - Bubbling through **limewater** turns solution cloudy
- Hydrogen** - Squeaky pop from it burning quick with O₂ to form **H₂O**

Anion tests:

- Carbonates** - (contain CO₃²⁻ ions) add **dilute acid** and connect to **lime water** if it's carbonate (water turns cloudy and carbon dioxide will form)
- Sulfates** - (contain SO₄²⁻) add **dilute hydrochloric acid** **precipitate**
- Halides** - Add **nitric acid** then **silver nitrate**. These are halide
Chloride = white precipitate **bromide** = cream precipitate
Iodide = yellow precipitate.

CATION TESTS:

Flame tests for metal ions:

- 1) Clean platinum wire in **dilute HCl**, hold in **blue flame**
 - 2) Dip loop into substance and hold in flame record colour.
- Lithium ions** = red flame **Calcium ions** = orange flame
 - Sodium ions** = yellow flame **Copper ions** = green flame
 - Potassium ions** = lilac flame
- Metal hydroxide tests:** form precipitates add sodium hydroxide
- Calcium (Ca²⁺)** = white precipitate **Copper (Cu²⁺)** = blue
 - Iron II (Fe²⁺)** = green **Iron III (Fe³⁺)** = brown
 - Aluminium (Al³⁺)** = white, then dissolves to colourless
 - Magnesium (Mg²⁺)** = white.

Chemistry Review Sheet

T. W. Graham Solomons



Chemestry Review Sheet:

Decoding **Chemestry Review Sheet**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chemestry Review Sheet**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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