

Chem 160: National Pi Day Activity

Mapping Atomic Structure (ChemFax by Flinn Sci.) (10 pts Lab extra credit)

Introduction: The mass of an atom is concentrated in its nucleus – a small, dense sphere with a diameter of about 1/100,000 that of the atom. Most of the size of an atom is apparently empty space! How can we translate the dimensions of the nucleus and the atom to things we can see and understand? In this activity, we will use a basketball and a map of the city to draw a scale model of the county.

Materials:

1 regulation-size men's basketball (or dimensions thereof)
Metric ruler, tape measure, compass or other method to draw a circle.

Procedure: (Show all of your calculations below.)

1. Using a tape measure, measure the circumference of the basketball in cm to at least 3 significant figures. (Or, if you do not have access to a basketball, look up the circumference on the internet! You may need to convert the measurement to cm.)

$$\text{Circumference of Basketball} = \underline{74.9 \text{ cm}} \quad 29.5 \text{ in} \times \frac{2.54 \text{ cm}}{\text{in}} = 74.9 \text{ cm}$$

2. The formula for the circumference of a sphere is: $C = 2\pi r$. Rearrange the formula to solve for the radius (r) and calculate the radius of the basketball in cm to three sig. figs. $4.7''$

$$\text{Radius of basketball} = \underline{11.9 \text{ cm}} \quad \begin{aligned} 74.9 &= 2\pi r \\ r &= \frac{74.9}{2\pi} = 11.92 \text{ cm} \end{aligned}$$

3. Imagine that the basketball represents the nucleus of an atom. Assuming that the radius of the entire atom is about 100,000 times larger than that of just the nucleus, calculate the radius of an atom that would have a nucleus the size of the basketball. Record the radius of this humongous atom in km to 3 sig. figs.

Calculations: (Box or circle your final calculation of the size of the radius of your atom.)

$$\begin{aligned} \text{Radius of Atom} &= 100,000 \times 11.9 \text{ cm} = 1,190,000 \text{ cm} \\ 1,190,000 \text{ cm} &\times \frac{1 \text{ m}}{100 \text{ cm}} \times \frac{1 \text{ km}}{1000 \text{ m}} = \boxed{11.9 \text{ km}} \end{aligned}$$

4. On the attached map, the "X" marks the approximate location of Solano College. Using a compass or some other method, draw a neat circle (not just sketched) on the map to illustrate the size the "humongous" atom would be if the nucleus was at SCC, and the nucleus were the size of the basketball.

Chemfax Lab Answers March Madness

Mark Steyvers



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