# Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)

 $C_3H_8 + O_2 - CO_2 + H_2O$ 

- a) If you start with 14.8 g of C<sub>3</sub>H<sub>8</sub> and 3.44 g of O<sub>2</sub>, determine the limiting reactant.
- b) Determine the number of moles of carbon dioxide produced.
- c) Determine the number of grams of H<sub>2</sub>O produced.
- d) Determine the number of grams of excess reactant left.
- Given the following equation:

Al2(SO3)3 + 6NaOH -----> 3Na2SO3 + 2Al(OH)3

- a) If 10.0 g of Al<sub>2</sub>(SO<sub>2</sub>)<sub>3</sub> is reacted with 10.0 g of NaOH, determine the limiting reactant.
- b) Determine the number of moles of Al(OH)<sub>3</sub> produced.
- c) Determine the number of grams of Na<sub>2</sub>SO<sub>3</sub> produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- 3. Given the following equation:

 $Al_2O_3$  + Fe ---->  $Fe_3O_4$  + Al

- a) If 25.4 g of Al<sub>2</sub>O<sub>3</sub> is reacted with 10.2 g of Fe, determine the limiting reactant.
- b) Determine the number of moles of Al produced.
- c) Determine the number of grams of Fe<sub>3</sub>O<sub>4</sub> produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- Consider the reaction

 $I_2O_5(g) + 5 CO(g) -----> 5 CO_2(g) + I_2(g)$ 

- a) 80.0 grams of I<sub>2</sub>O<sub>5</sub>, reacts with 28.0 grams of carbon monoxide, CO. Determine the mass of iodine I<sub>2</sub>, which could be produced?
- b) If, in the above situation, only 0.160 moles, of iodine, I<sub>2</sub> was produced.
  - i) What mass of iodine was produced?
  - ii) What percentage yield of iodine was produced?
- 5. Zinc and sulfur react to form zinc sulfide according to the equation

Zn + S -----> ZnS

If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- a) Which chemical is the limiting reactant?
- b) How many grams of ZnS will be formed?
- c) How many grams of the excess reactant will remain after the reaction is over?
- 6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?
- 7. How many grams of Al<sub>2</sub>S<sub>3</sub> are formed when 5.00 grams of Al is heated with 10.0 grams S?

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**Thomas Griffiths** 

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- 1. Understanding the eBook Chemfax Limiting Reagent Lab Answers
  - The Rise of Digital Reading Chemfax Limiting Reagent Lab Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chemfax Limiting Reagent Lab Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemfax Limiting Reagent Lab Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chemfax Limiting Reagent Lab Answers
  - Personalized Recommendations
  - Chemfax Limiting Reagent Lab Answers User Reviews and Ratings
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- 8. Staying Engaged with Chemfax Limiting Reagent Lab Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemfax Limiting Reagent Lab Answers
- 9. Balancing eBooks and Physical Books Chemfax Limiting Reagent Lab Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemfax Limiting Reagent Lab Answers
- 10. Overcoming Reading Challenges
  - o Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemfax Limiting Reagent Lab Answers
  - Setting Reading Goals Chemfax Limiting Reagent Lab Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemfax Limiting Reagent Lab Answers
  - Fact-Checking eBook Content of Chemfax Limiting Reagent Lab Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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