

Name: Answer Key

Period: \_\_\_\_\_

## Titration Practice Worksheet

Find the requested quantities in the following problems:

- 1) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?

$$[HCl] = 0.0432 \text{ M}$$

- 2) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

$$[NaOH] = 0.0036 \text{ M}$$

- 3) If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulfuric acid solution ( $H_2SO_4$ ), what is the concentration of the  $H_2SO_4$  solution?

✓ - 2 hydrogens in this acid: diprotic acid

$$[H_2SO_4] = \frac{0.2 \text{ M}}{2} = 0.1 \text{ M}$$

- 4) Can I titrate a solution of unknown concentration with another solution of unknown concentration and still get a meaningful answer? Explain your answer in a few sentences.

No

- 5) Explain the difference between an endpoint and equivalence point in a titration.

endpoint: lab determined, indicator color change

equivalence: moles of  $H^+$  (acid) = moles of  $OH^-$  (base)

# Chemfiesta Acid And Base Practice Answers

**Izaak Maurits Kolthoff**



## Chemfiesta Acid And Base Practice Answers:

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*The ABC of Acid-base Chemistry ...* Horace Willard Davenport, 1974

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