

# LIMITING REACTANT STOICHIOMETRY WORKSHEET

Name _____	Section _____	Date _____
<b>Limiting Reactants</b>		
<b>Directions:</b> Determine the limiting reactant in each problem.		
1) In $2\text{NaOH} + \text{MgCl}_2 \rightarrow \text{Mg(OH)}_2 + 2\text{NaCl}$ , 5.818 grams of NaOH react with 2.712 grams of $\text{MgCl}_2$ .		
2) 0.0718 grams of HCl react with 0.0002 grams of $\text{Ca(OH)}_2$ in the reaction $2\text{HCl} + \text{Ca(OH)}_2 \rightarrow 2\text{H}_2\text{O} + \text{CaCl}_2$ .		
3) 0.2000 grams of Nitrogen in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ combine with 0.2008 grams of Hydrogen.		

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3) 0.2000 grams of Nitrogen in the reaction $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ combine with 0.2008 grams of Hydrogen.		
4) In the reaction $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$ , 10.000 grams of $\text{O}_2$ react with 0.908 grams of $\text{C}_2\text{H}_2$ .		
5) In the reaction $2\text{Al(OH)}_3 + 3\text{H}_2\text{CO}_3 \rightarrow \text{Al}_2(\text{CO}_3)_3 + 6\text{H}_2\text{O}$ , 1.18 grams of $\text{Al(OH)}_3$ and 1.012 grams of $\text{H}_2\text{CO}_3$ combine.		
6) 0.838 grams of sodium and 0.7021 grams of sodium nitrate react in the following process: $10\text{Na} + 2\text{NaNO}_3 \rightarrow 6\text{Na}_2\text{O} + \text{N}_2$ .		
7) 0.0700 grams of $\text{AlBr}_3$ react with 1.201 grams of $\text{H}_2\text{SO}_4$ to produce products in the following reaction: $2\text{AlBr}_3 + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 6\text{HBr}$ .		
8) When $\text{K}_3\text{PO}_4$ and NaOH react in $\text{K}_3\text{PO}_4 + 3\text{NaOH} \rightarrow 3\text{KOH} + \text{Na}_3\text{PO}_4$ , 5.41 grams of NaOH react with 3.88 grams $\text{K}_3\text{PO}_4$ .		
9) In $\text{Pt}_2(\text{NO})_2 + \text{H}_2\text{S} \rightarrow \text{PtS} + \text{HNO}_3$ , 10.67 grams of $\text{Pt}_2(\text{NO})_2$ react with 7.98 grams of $\text{H}_2\text{S}$ .		
10) 0.4000 grams of $\text{H}_2\text{SO}_4$ combine with 0.2480 grams of $\text{Ba(OH)}_2$ in the reaction $2\text{H}_2\text{SO}_4 + \text{Ba(OH)}_2 \rightarrow \text{BaSO}_4 + 2\text{H}_2\text{O}$ .		

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A CHEMISTRY WORKSHEET  
10 PROBLEMS  
WORKED ANSWER KEY

# Chemfiesta Stoichiometry Limiting Reagents Practice Answers

**Cosmas C Mbuzi**



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Calculating in Chemistry Cosmas C Mbuzi, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to covert various quantities to moles or amount of substance This particularly useful to a beginner Chapter 2 describes how to calculate reacting reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the

calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

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