

pH and Buffers

ACIDS AND BASES

Bronsted and Lowry have defined acids as proton donors and bases as proton acceptors. Each acid, therefore, has always a conjugated base.

Acid → Base + H⁺

 $HCI \rightarrow CI^- + H^+$

Alkali refers for the compound that yields hydroxyl ions on dissociation.

Base → Conjugate Acid + OH

 $NaOH \rightarrow Na^{+} + OH^{-}$

The strength of acids or bases is the ability to lose H* or OH*. Weak acids and bases dissociate only to limited extent.

 $H_2SO_4 \rightarrow H^* + HSO_4$

 $KOH \rightarrow K^* + OH^-$

Potassium hydroxide

 $CH_3COOH \rightarrow CH_3COO^- + H^-$

(Weak acid)

 $C_6H_5NH_2 + H^+ \rightarrow C_6H_5NH_3^+$

(Weak base) (Aniline)

Chemical Basis Of Life Module A Anchor 2 Answers

John Desmond Bernal

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