

coefficient of 2 in front of O₂, then we will increase the number of oxygen atoms in the reactants by 2. The new equation is:
$$\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$$

When we check the number of atoms again, we find that the number of atoms of each element in the reactants is the same as the number in the products. The equation is now balanced.

Example 3: Balancing chemical equations 3

Question

In our bodies, sugar (C₆H₁₂O₆) reacts with the oxygen we breathe in to produce carbon dioxide, water and energy. Write the balanced equation for this reaction.

Answer

Identify the reactants and products in the reaction.

Reactants: sugar (C₆H₁₂O₆) and oxygen (O₂)

Products: carbon dioxide (CO₂) and water (H₂O)

Write the equation



Count the number of atoms of each element in the reactants and in the products

Reactants: C=6; H=12; O=8

Products: C=1; H=2; O=3

Balance the equation

It is easier to start with carbon as it only appears once on each side. If we add a 6 in front of CO₂, the equation looks like this:



Reactants: C=6; H=12; O=8

Products: C=6; H=2; O=13

Change the coefficients again to try to balance the equation.

Let us try to get the number of hydrogens the same this time.



Reactants: C=6; H=12; O=8

Products: C=6; H=12; O=18

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Julie R. Frentrup



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