

Chemical reactions involve changing one substance into another substance by rearranging atoms. However, during a chemical reaction atoms of one element cannot change into atoms of another element. The reason this change cannot occur is that chemical reactions only involve an atom's electrons – the nucleus remains unchanged.

Recall that an atom's identity is based on its number of protons. Since protons are in the nucleus and chemical reactions do not involve the nucleus, the atom remains unchanged. However, there are some reactions that do involve changes in the nucleus. These are called nuclear reactions and do change one atom of an element into an atom of a different element.

1. Fill in the table below as a review. **You will need your periodic table for this!** Remember the atomic number (or # of protons) determines the element. If you have four protons and seven neutrons you have beryllium. The same is true if you have four protons and six neutrons...you still have beryllium.

Isotope	Total Protons (Atomic #)	Total Neutrons (Mass # - Atomic #)	Mass Number*	Total Electrons Outside Nucleus	Format for Nuclear Equation
K-40	19	21	40	19	${}^{40}_{19}\text{K}$
Li-6	3	3	6	3	${}^6_3\text{Li}$
He-3	2	1	3	2	${}^3_2\text{He}$
I-131	53	78	131	53	${}^{131}_{53}\text{I}$
Sr-90	38	52	90	38	${}^{90}_{38}\text{Sr}$

\*NOTE: Do NOT use the mass numbers from your periodic table.

Radioactivity is when a substance spontaneously emits radiation. Radioactive atoms (or radioisotopes) emit radiation because their nuclei are unstable. Unstable nuclei lose energy by emitting radiation in a spontaneous process called radioactive decay. Unstable radioactive atoms undergo radioactive decay until they form stable nonradioactive atoms. There are several types of radiation emitted during radioactive decay.

#### Types of Radiation: Alpha, Beta, and Gamma

Three types of radiation have been discovered. The types are called alpha, beta and gamma. **Alpha** rays turned out to be small particles of matter with a charge of +2 and a mass of 4 amu. It has been proved that an alpha particle contains two protons and two neutrons – it is identical to the nucleus of a helium atom. In fact, when an alpha particle slows down and gains two electrons it becomes a helium atom. The Greek letter alpha ( $\alpha$ ) is used to represent this particle but in equations to keep track of mass and protons we must use  ${}^4_2\text{He}$ . **Betas** were also found to be particles; they are simply high speed electrons. We use the Greek letter beta ( $\beta$ ), but in equations  ${}^0_{-1}e$  is used. When a beta slows down it becomes an electron. **Gamma** rays ( $\gamma$ ) are not particles; they are high energy electromagnetic radiation. They are photons (light) with no charge or mass so we simply write  ${}^0_0\gamma$  in our equations.

Example 1: Thorium-232 decays by emitting an alpha and a gamma.



# Chemical Nuclear Packet Worksheet 1 Radioactivity

**Gregory R. Choppin**



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