# Periodic Table WS 8

#### Chemical Periodicity Review Sheet

#### H/Chemistry

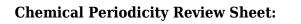
#### A. Completion

Use this completion exercise to check your knowledge of the terms and your understanding of the concepts introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

The periodic table	organizes the elements into vertical	and horizonta	ı
in order of increasing	r of increasing The table is constructed so that elements that have		nilar chemical
properties are in the same_	. The elements in Groups 17	A through 7A (or 1,2, 13-1	7) are called the
	The make up Group	18. The elements in Grou	ps 2 and 13 are
interrupted in periods 4 through 7 by the		and in periods 6 and 7 by the	
The atoms of the ne	oble gas elements have their outermost s	and sublevels filled.	The outermost s
and p sublevels of the repre	sentative elements are		
Atomic radii genera	ally as you move from lef	t to right in a period. Ator	nic radius
nerally within a given group because there are more		re	occupied and an
increased	_ effect, despite an increase in nuclear _		
The energy required to remove an electron from an atom is known as the			This
quantity generally	as you move left to right across a period, and		as you
move down a group. Its tre	nd is the opposite to that of		
The ability of a bor	ded atom to attract electrons to itself is	known as	, and this
quantity	as we move from left to right acr	oss a period, and	as w
move down a group.			

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