

Periodic Table WS 8

Chemical Periodicity Review Sheet

H/Chemistry

A. Completion

Use this completion exercise to check your knowledge of the terms and your understanding of the concepts introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

The periodic table organizes the elements into vertical _____ and horizontal _____ in order of increasing _____. The table is constructed so that elements that have similar chemical properties are in the same _____. The elements in Groups 1A through 7A (or 1,2, 13-17) are called the _____. The _____ make up Group 18. The elements in Groups 2 and 13 are interrupted in periods 4 through 7 by the _____ and in periods 6 and 7 by the _____.

The atoms of the noble gas elements have their outermost s and _____ sublevels filled. The outermost s and p sublevels of the representative elements are _____.

Atomic radii generally _____ as you move from left to right in a period. Atomic radius generally _____ within a given group because there are more _____ occupied and an increased _____ effect, despite an increase in nuclear _____.

The energy required to remove an electron from an atom is known as the _____. This quantity generally _____ as you move left to right across a period, and _____ as you move down a group. Its trend is the opposite to that of _____.

The ability of a bonded atom to attract electrons to itself is known as _____, and this quantity _____ as we move from left to right across a period, and _____ as we move down a group.

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