

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $\text{PCl}_5$ )
  - b. uranium hexafluoride ( $\text{UF}_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $\text{KMnO}_4$
  - b.  $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $\text{O}_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $\text{C}_6\text{H}_{12}\text{O}_6$
  - b.  $\text{NaHCO}_3$
  - c.  $\text{C}_2\text{H}_2$
  - d.  $\text{KNH}_4\text{SO}_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $\text{PbO}$ )
  - b. 0.75 mol hydrogen sulfide ( $\text{H}_2\text{S}$ )
  - c. 0.00100 mol silicon tetrahydride ( $\text{SiH}_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $\text{O}_2$ )
  - e. 2.30 mol ethylene glycol ( $\text{C}_2\text{H}_6\text{O}_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $\text{N}_2\text{O}_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $\text{Ca}_3(\text{PO}_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $\text{NaBr}$ )
  - b. 13.5 g magnesium nitrate [ $\text{Mg}(\text{NO}_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $\text{MgCl}_2$ )
  - d. 0.00100 g monochloromethane ( $\text{CH}_3\text{Cl}$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $\text{C}_3\text{H}_8(\text{OH})_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $\text{CH}_4$ ) gas?
9. How many moles is 135 L of ammonia ( $\text{NH}_3$ ) gas at STP?

# Chemical Quantities Practice Problem Answers

**Stanley Wallace Morse**



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*A Visual Analogy Guide to Chemistry, 2e* Paul A Krieger, 2018-02-01 A Visual Analogy Guide to Chemistry is the latest in the innovative and widely used series of books by Paul Krieger This study guide delivers a big picture view of difficult concepts and effective study tools to help students learn and understand the details of general organic and biochemistry topics A Visual Analogy Guide to Chemistry is a worthwhile investment for any introductory chemistry student

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Advances in Computer Assisted Learning P. R. Smith, 2014-05-23 Advances in Computer Assisted Learning contains selected proceedings from the CAL Symposium on Computer Assisted Learning held at the University of Nottingham in the UK in 1985 This book reviews advances in computer assisted learning in the areas of curriculum development visually handicapped and disabled students project work in schools television viewdata and video applications database applications and engineering education and training This monograph has 35 chapters and opens with a discussion on the computing aspects of interactive video focusing on the design and production of the software used to

control the videodisc developed by the Open University in the UK The next chapter illustrates a variety of case studies whereby local viewdata has been exploited by both teachers and their pupils in different parts of Europe Attention then turns to the use of computer assisted communication in the education of the visually impaired the use of microcomputers in teaching electronics and theoretical considerations in selecting software for language arts This text will be of interest to educators and policymakers who want to implement computer technology in the classroom *Chemistry: Matter & Change, Solving Problems - A Chemistry Handbook* McGraw Hill, 2001-08 *Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change* Chemical Symbolism and Calculations Stanley Wallace Morse, 1953 *General Organic and Biological Chemistry* Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features **Introduction to Chemistry** Richard C. Bauer, James P. Birk, Pamela Marks, 2024 The sixth edition of *Introduction to Chemistry* continues to build on our belief that students learn best when the text and our classroom presentations focus on a conceptual approach to chemistry Our class meetings are significantly different from traditional lecture presentations in many ways Beginning with the first week of classes and continuing through the rest of the semester we follow a sequence of topics that allows us to explain macroscopic phenomena from a molecular perspective This approach places emphasis on conceptual understanding over algorithmic problem solving

**The Chemical News and Journal of Physical Science**, 1907 **Crystal Structures** M Ladd, 1999-06 This book presents and discusses those common crystal structures that would be encountered by students taking chemistry or any subject within which chemistry forms a significant component With many worked examples and a wide selection of problems with solutions Includes instructions for making simple stereoviewers and computer programs in a thorough treatment of binary alloys and three dimensional packing in molecular solids *The Calculations of General Chemistry, with Definitions, Explanations, and Problems* William Jay Hale, 1920 Chemical News and Journal of Physical Science William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1888 **Chemistry, Problem-Solving Worktext** James E. Brady, Frederick Senese, 2002-06-13 The images on the cover call attention to the relationship between macro observations and the intimate structure of chemical substances and the changes both chemical and physical that they undergo Fireworks One of the ingredients is phosphorus a molecular form of which is believed to consist of linked tetrahedra of phosphorus atoms The chemical reaction of phosphorus with oxygen is partly responsible for the spectacular show of light Carbon The element is found in several forms including the familiar diamond and another recently discovered sooty substance that consists of soccer ball shaped molecules often referred to as buckeyballs Diamond is not the most stable form of carbon and

is created from other forms of carbon at high temperatures and pressures deep within the earth Acetylene torch Cutting steel is possible because of the intense heat generated by the chemical reaction of acetylene with oxygen a reaction between molecules of  $C_2H_2$  and  $O_2$  to give  $CO_2$  and  $H_2O$  Hot air balloon The air that helps it rise is heated by the combustion of molecules of propane each composed of three carbon and eight hydrogen atoms Stormy weather The evaporation of water serves to store energy provided by the sun Subsequent condensation of the water vapor releases this energy and is the basis of all the weather systems on our planet Watts' Dictionary of Chemistry Henry Watts, Henry Forster Morley, Matthew Moncrieff Pattison Muir, 1894 **Chemicals**, 1925 **The Industrial Chemist and Chemical Manufacturer**, 1929

*Chemistry* Neil D. Jespersen, Alison Hyslop, 2021-12-03 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills with a focus on integrating metacognition and three dimensional learning into the text When students know what they know they are better able to learn and incorporate the material Providing a total solution through New WileyPLUS by fully integrating the enhanced etext with online assessment answer specific responses and additional practice resources The 8th edition continues to emphasize the importance of applying concepts to problem solving to achieve high level learning and increase retention of chemistry knowledge Problems are arranged in an intuitive confidence building order State board questions and answers Rudolph Max Goepp, 1913 **Chemical, Color and Oil Record**, 1925 **Basic Mathematics for Beginning Chemistry** Dorothy M. Goldish, 1979

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